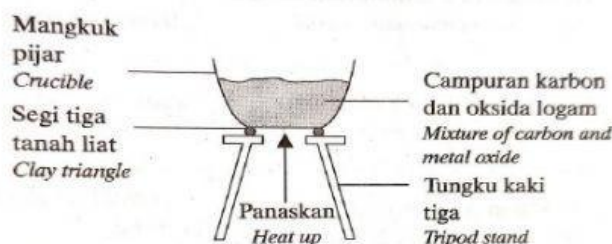


Name: _____

Class: _____

4.2 Reactivity series of metals

1. Experiment below to determine the position of carbon in reactivity series of metal through heating of the different substances such as zinc oxide, aluminium oxide and iron oxide.



a) Complete the word equation for each metal oxide reaction with carbon, if any.

(i) Zinc oxide + Carbon \longrightarrow

(ii) Aluminium oxide + Carbon \longrightarrow

(iii) Iron oxide + Carbon \longrightarrow

b) State metals that are less reactive than carbon. Explain your answer.

Reason : Oxides of metals which are less reactive than carbon will _____ by carbon and turn into the _____.

c) Complete the sequence of metal arrangement according to increasing reactivity towards oxygen.

\longrightarrow \longrightarrow \longrightarrow

$\xrightarrow{\hspace{15em}}$ Increasing reactivity

2. Complete the blanks with correct answers:

Aluminium	Ferrum	Oxygen	Zinc	Lead	hydrogen	Carbon
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a) State the metal that reacts with hydrogen.

b) Hydrogen cannot remove oxygen from the oxide and oxide.

c) Copper oxide + \longrightarrow Copper + Water

d) can reduce zinc oxide into zinc.

e) Lead + \longrightarrow Lead oxide

