INTRODUCTION TO REDOX CHEMISTRY

Complete the sentences below:

1. Oxidation occurs when a substance _____ electrons.
2. Reduction occurs when a substance _____ electrons.

Determine and then write the oxidation state of the underlined element in the space provided. You must include "+" or "-" sign in front of the oxidation number.

Example:

$\underline{\mathbf{Cr}}_2O_7^{2-}$ Oxidation State: +6

3. $\underline{\mathbf{P}}_2O_{10}$ Oxidation State:_____
4. $\underline{\mathbf{N}}_2O$ Oxidation State:_____
5. $\underline{\mathbf{S}}O_3$ Oxidation State:_____
6. $Na_2\underline{\mathbf{S}}O_3$ Oxidation State:_____
7. $\underline{\mathbf{Cr}}O_4^{2-}$ Oxidation State:_____
8. $\underline{\mathbf{V}}O_3^-$ Oxidation State:_____
9. $\underline{\mathbf{C}}O$ Oxidation State:_____
10. $\underline{\mathbf{Sn}}O_2$ Oxidation State:_____