



Pareeksha Kirana – 2021

From Team Vijnana Kirana

CHEMISTRY

Ch : 5 PERIODIC CLASSIFICATION OF ELEMENTS -2

- Which of the following is the correct increasing order of the atomic radii of the elements oxygen, fluorine and nitrogen?
a) $O < F < N$ b) $N < F < O$ c) $O < N < F$ d) $F < O < N$
- The elements A, B and C belong to group 2, 14 and 16 respectively, of the periodic table. Which of the elements will form covalent bonds?
a) A and B b) B and C c) C and A d) A
- An element X belongs to the 3rd period and 1st group of the periodic table. What is the number of valence electrons in its atom?
a) 1 b) 3 c) 6 d) 8
- An element M is in group 13th of the periodic table, the formula of its oxide is
a) MO b) M_2O_3 c) M_3O_2 d) MO_2
- Observe the table and identify the formula of oxide of lithium

2nd period	Li	Be	B	C	N	O	F	Ne
Valency	-1	2	3	4	3	2	1	0

a) LiO b) Li_2O c) LiO_2 d) Li_2O_3
- In the second period the most metallic element is
a) Be - as it has more protons b) Li - it is with least nuclear charge
c) F- as it is most electro negative d) Ne – as it has completely filled shell
- Electronic configuration of carbon is
a) 2, 2 b) 2, 3 c) 2, 4 d) 2, 5
- Number of valence electrons found in the element with atomic number 19 is
a) 1 b) 2 c) 3 d) 4
- A, B, C, D, E are the elements belongs to group 1, 2, 13, 14, 16 respectively. Most electronegative element in the group is
a) A b) D c) B d) E
- Correct arrangement of the elements according to increasing order of their nuclear charge is
a) $Li > Be > B > C > N > O > F$ b) $Li < Be < B < C < N < O < F$
c) $Li > B > Be > C > O > N > F$ d) $Li < Be < B < C < N < O < F$
- Identify the element that has the tendency to lose the electrons most easily.
a) Ca b) Na c) K d) Mg

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05-06-21

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