

3. Substance: Carbon dioxide

Phase: Solid

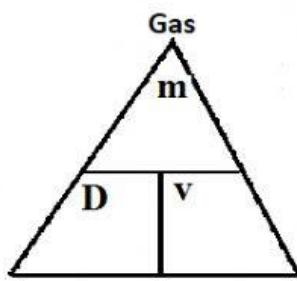
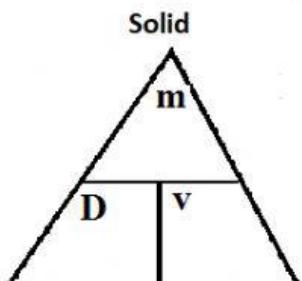
Volume Mass
100 cm³ 156 g

Phase: Gas
Density Volume
_____ g/cm³ 100 cm³

Mass
0.198 g
4 decimals

Density _____ g/cm³
answer
5 decimals

Name of phase change from solid to gas _____



Which is more dense, the solid or gas? **solid** **gas**

How does the arrangement of the particles in a solid and a gas explain your answer?

The particles in a **solid** **gas** are packed more closely together this **increases** **decreases** density.

4. Substance: Mercury

Phase: Solid

Volume Mass
10 cm³ 136.5 g

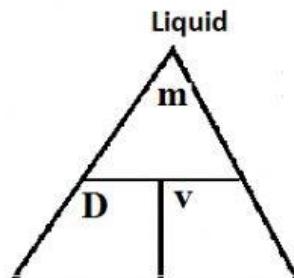
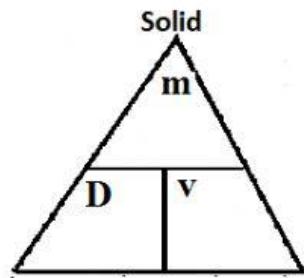
Density _____ g/cm³
answer
2 decimals

Phase: Liquid

Volume Mass
10 cm³ 135.9 g

Density _____ g/cm³
answer
2 decimals

Name of phase change from liquid to solid _____



Which is more dense, the solid or liquid? **solid** **liquid**

How does the arrangement of the particles in a solid and a liquid explain your answer?

The particles in a **solid** **liquid** are packed more closely together this **increases** **decreases** density.

5. Substance: Pentane

Phase: Liquid

Volume Mass
3800 cm³ 2370 g

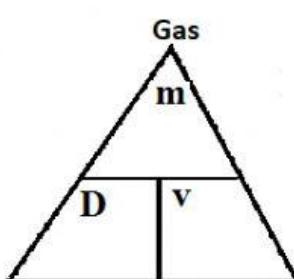
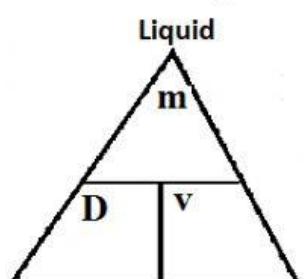
Density _____ g/cm³
answer
2 decimals

Phase: Gas

Volume Mass
3800 cm³ 10.7 g

Density _____ g/cm³
answer
4 decimals

Name of phase change from gas to liquid _____



Which is more dense, the liquid or gas? **liquid** **gas**

How does the arrangement of the particles in a liquid and a gas explain your answer?

The particles in a **liquid** **gas** are packed more closely together this **increases** **decreases** density.