

3. Substance: Carbon dioxide

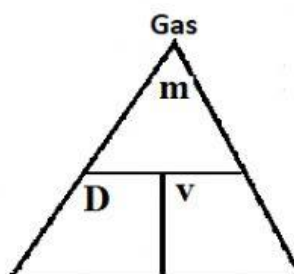
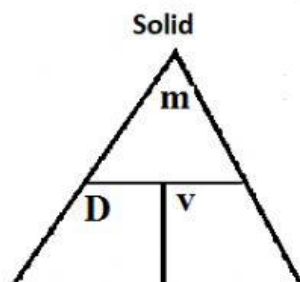
Phase: Solid

Volume	Mass	Density
100 cm ³	156 g	<u> </u> g/cm ³
		answer 2 decimals

Phase: Gas

Volume	Mass	Density
100 cm ³	0.198 g	<u> </u> g/cm ³
		answer 5 decimals

Name of phase change from solid to gas



Which is more dense, the solid or gas? **solid gas**

How does the arrangement of the particles in a solid and a gas explain your answer?

The particles in a **solid gas** are packed more closely together this **increases decreases** density.

4. Substance: Mercury

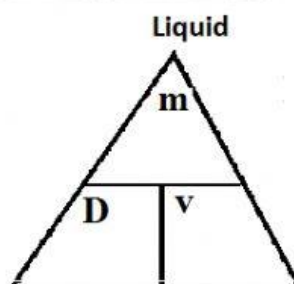
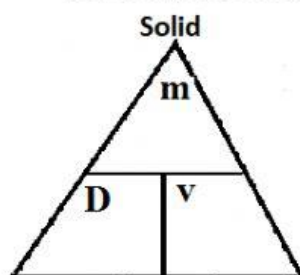
Phase: Solid

Volume	Mass	Density
10 cm ³	136.5 g	<u> </u> g/cm ³
		answer 2 decimals

Phase: Liquid

Volume	Mass	Density
10 cm ³	135.9 g	<u> </u> g/cm ³
		answer 2 decimals

Name of phase change from liquid to solid



Which is more dense, the solid or liquid? **solid liquid**

How does the arrangement of the particles in a solid and a liquid explain your answer?

The particles in a **solid liquid** are packed more closely together this **increases decreases** density.

5. Substance: Pentane

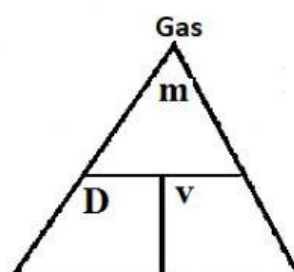
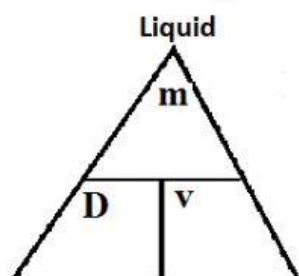
Phase: Liquid

Volume	Mass	Density
3800 cm ³	2370 g	<u> </u> g/cm ³
		answer 2 decimals

Phase: Gas

Volume	Mass	Density
3800 cm ³	10.7 g	<u> </u> g/cm ³
		answer 4 decimals

Name of phase change from gas to liquid



Which is more dense, the liquid or gas? **liquid gas**

How does the arrangement of the particles in a liquid and a gas explain your answer?

The particles in a **liquid gas** are packed more closely together this **increases decreases** density.