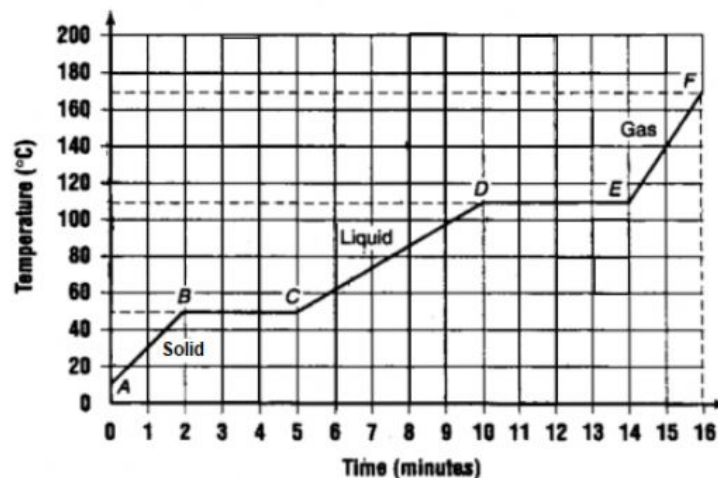


# Kinetic Theory

**Directions:** Look carefully at the graph. It was drawn from the data collected when a substance was heated at a constant rate. To heat at a constant rate means to add heat evenly as time passes. Use the graph to complete the paragraphs that follow.



At the start of observations, Point A, the substance exists in the 1. \_\_\_\_ state. The temperature at this point is 2. \_\_\_\_\_. As energy is absorbed, the temperature of the substance rises at a constant rate for two minutes. At Point B, the temperature is 3. \_\_\_\_\_, and the solid begins to 4. \_\_\_\_\_. The temperature remains constant until the change from solid to 5. \_\_\_\_\_ is complete. It has taken three minutes to add enough energy to melt the solid completely. From Point C to Point D, the substance is in the 6. \_\_\_\_\_ state. Its temperature rises at a constant rate to 7. \_\_\_\_\_. The temperature remains constant while the liquid changes to a 8. \_\_\_\_\_. At Point E, the substance exists as a 9. \_\_\_\_\_. Its temperature rises evenly as energy is added.

When the gaseous substance is allowed to cool, it releases energy. The cooling curve will be the reverse of the warming curve. Energy will be released as the substance changes from a 10. \_\_\_\_\_ to a 11. \_\_\_\_\_ and also from a 12. \_\_\_\_\_ to a 13. \_\_\_\_\_. The amount of energy released during condensation will be the same as the amount absorbed during vaporization.