PERCENT YIELD PROBLEMS

(Take your time, use the periodic table I've provided on the STREAM to calculate molar mass, watch the videos provided for guidance)

1. If, in the reaction below 32 grams of C_2H_6 and 39 grams of O_2 produces 28 grams of CO_2, what is the % yield?

$$\mathbf{2C_2H_6 + 7O_2 \rightarrow 4CO_2 + 6H_2O}$$

2. If, in the reaction below 32 grams of C_2H_6 produces 44 grams of CO_2, what is the % yield?

$$\mathbf{2C_2H_6 + 7O_2 \rightarrow 4CO_2 + 6H_2O}$$

3. If, in the reaction below, 80 grams of Cl_2 produces 38 grams of CCl_4 what is the % yield?

$$\mathbf{CS_2 + 3Cl_2 \rightarrow CCl_4 + S_2Cl_2}$$

4. If, in the reaction below, 4 grams of H_2O produces 0.67 grams of HF what is the % yield?

$$\mathbf{CH_3COF + H_2O \rightarrow CH_3COOH + HF}$$

LIVEWORKSHEETS