

Assessment.

Fill in the blanks and submit the answers to teacher code ThTDSNuqzs

1. Reacting a solid with an acid to make soluble salts can involve an acid and a metal, but only for moderately reactive metals from magnesium to _____ in the reactivity series.
2. When preparing copper(II) sulfate crystals, the solid copper(II) oxide added to the beaker is _____ in color.
3. In the preparation of copper(II) sulfate, having unreacted copper(II) oxide left at the bottom of the beaker indicates that all the acid has been _____.
4. For safety reasons, direct heating of the acid with a Bunsen burner is not recommended; instead, it is best heated in a boiling tube placed in a _____.
5. Evaporating a copper(II) sulfate solution completely to dryness will not produce blue crystals, but will instead form a white powder of _____ copper(II) sulfate.
6. When many salts form crystals, water from the solution becomes chemically bound up with the salt, which is called water of _____.
7. A salt that contains water of crystallisation within its structure is said to be _____.
8. When making magnesium sulfate crystals by adding magnesium to sulfuric acid, the reaction mixture fizzes because _____ gas is given off.
9. Unlike the reaction with copper(II) oxide, when adding magnesium to dilute sulfuric acid, the acid does not have to be _____.
10. A laboratory technician can check if a concentrated salt solution is close to being saturated by dipping a _____ into it to see if crystals form upon removal.