

1. A substance which oxidises itself and reduces other is known as -
- (a) oxidising agent (b) reducing agent
(c) both of these (d) none of these
2. Which of the following reactions involves the combination of two elements?
- (a) $\text{CaI} + \text{CO}_2 \rightarrow \text{CaCO}_3$ (b) $4\text{Na} + \text{O}_2 \rightarrow 2\text{Na}_2\text{O}$
(c) $\text{SO}_2 + \frac{1}{2}\text{O}_2 \rightarrow \text{SO}_3$ (d) $\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4\text{Cl}$
3. When hydrogen sulphide gas is passed through a blue solution of copper sulphate, a black precipitate of copper sulphide is obtained and the sulphuric acid so formed remains in the solution. The reaction is an example of-
- (a) a combination reaction
(b) a displacement reaction
(c) a decomposition reaction
(d) a double decomposition reaction
4. Which of the following is a physical change?
- (a) Formation of curd from milk
(b) Ripening of fruits
(c) Getting salt from sea water
(d) Burning of wood

- 5.** What happens when copper rod is dipped in iron sulphate solution?
- (a) Copper displaces iron
 - (b) Blue colour of copper sulphate solution is obtained
 - (c) No reaction takes place
 - (d) Reaction is exothermic
- 6.** A student added dilute HCl to a test tube containing zinc granules and made following observations :
- (a) the zinc surface became dull and black
 - (b) a gas evolved which burnt with a pop sound
 - (c) the solution remained colourless
 - (d) the solution becomes green in colour
- 7.** A dilute solution of sodium carbonate was added to two test tubes - one containing dil HCl (a) and the other containing dilute NaOH(b). The correct observation was-
- (a) a brown coloured gas liberated in test tube A
 - (b) a brown coloured gas liberated in test tube B
 - (c) a colourless gas liberated in test tube A
 - (d) a colourless gas liberated in test tube B
- 8.** A balanced chemical equation is in accordance with-
- (a) Avogadro's law
 - (b) law of multiple proportion
 - (c) law of conservation of mass
 - (d) law of gaseous volumes