Bellringers Week 8: 10/13 - 10/16

Day 1: Monday 10/13

- How many electrons can fit into each of these sublevels:

s: **p:** **d:** **f:**

- Use your periodic table blocks and tell me how many and which orbitals are on each of the following energy levels: for example (1, s)

Ring/Energy Level/Orbital 1:

Ring/Energy Level/Orbital 2:

Ring/Energy Level/Orbital 3:

Ring/Energy Level/Orbital 4:

Ring/Energy Level/Orbital 5:

Ring/Energy Level/Orbital 6:

Ring/Energy Level/Orbital 7:

Day 2: Tuesday 10/14

- Use shorthand electron configuration to write out the electron configurations for these elements:

$_{38}$Sr:

$_{33}$As:

$_{63}$Eu:

$_{74}$W:

Day 3: Wednesday 10/15

- PSAT Day

Day 4: Thursday 10/16

Match each rule to its definition:

Aufbau principle	orbitals can hold a maximum of two electrons with opposite spins
Hund's rule	Fill orbitals from lowest to highest energy
Pauli exclusion principle	electrons first occupy each orbital singly before any orbital is paired up

Day 5: Friday 10/17

- NO SCHOOL