

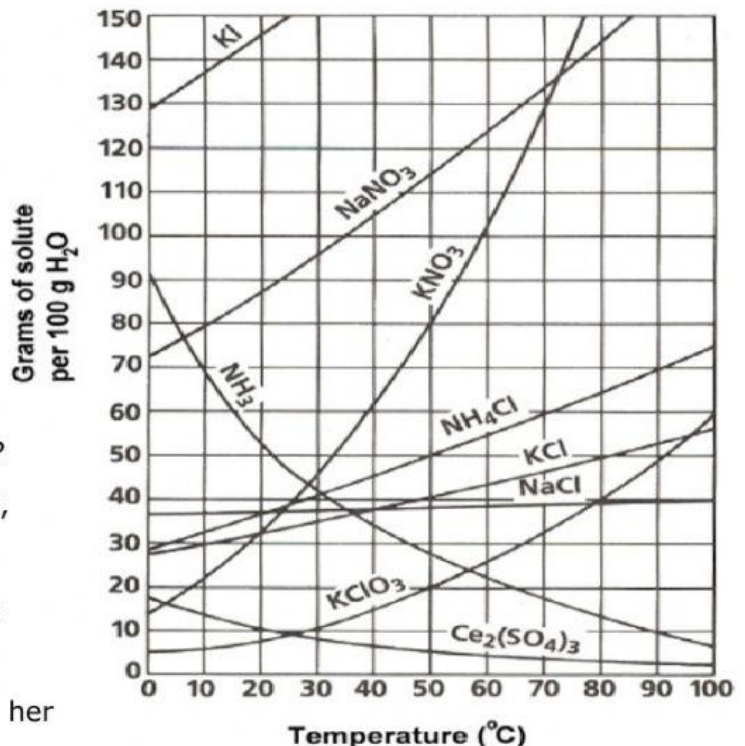
## Solubility Curve Practice

**Part 1: Click on the correct word that matches the definition:**

1. A solution where more solute can dissolve:
2. A solution where no more solute can be added. It has the maximum of amount of solute and some may settle at the bottom:
3. The solution is unstable and cannot hold any more solute. Crystals start to form at the bottom:

**Part 2: Find the solubility for each salt using the Solubility Curve. (make sure to use the unit: g/100g of water)**

1.  $\text{KNO}_3$  at  $50^\circ\text{C}$ =
2.  $\text{Ce}_2(\text{SO}_4)_3$  at  $20^\circ\text{C}$ =
3.  $\text{NaNO}_3$  at  $25^\circ\text{C}$ =



- 1) Which of the **ABOVE** 3 substances is most soluble in water?
- 2) Terry dissolves 129g of KI at  $0^\circ\text{C}$ , his solution is
- 3) Terry dissolves 120g of  $\text{NaNO}_3$  at  $60^\circ\text{C}$ , his solution is
- 4) Ellie dissolves 80g of KCl at  $90^\circ\text{C}$ , her solution is
- 5) Which compound is the least soluble at  $40^\circ\text{C}$ ?
- 6) What temp. would you have to heat 80g of  $\text{NaNO}_3$  solution to for it to be saturated?