

Solubility Curve Practice

Part 1: Click on the correct word that matches the definition:

1. A solution where more solute can dissolve:
2. A solution where no more solute can be added. It has the maximum of amount of solute and some may settle at the bottom:
3. The solution is unstable and cannot hold any more solute. Crystals start to form at the bottom:

Part 2: Find the solubility for each salt using the Solubility Curve. (make sure to use the unit: g/100g of water)

1. KNO_3 at 50°C =

2. $\text{Ce}_2(\text{SO}_4)_3$ at 20°C =

3. NaNO_3 at 25°C =

- 1) Which of the **ABOVE** 3 substances is most soluble in water?
- 2) Terry dissolves 129g of KI at 0°C , his solution is
- 3) Terry dissolves 120g of NaNO_3 at 60°C , his solution is
- 4) Ellie dissolves 80g of KCl at 90°C , her solution is
- 5) Which compound is the least soluble at 40°C ?
- 6) What temp. would you have to heat 80g of NaNO_3 solution to for it to be saturated?

