

# Weekly Warm Ups

Honors Chemistry

**Monday**

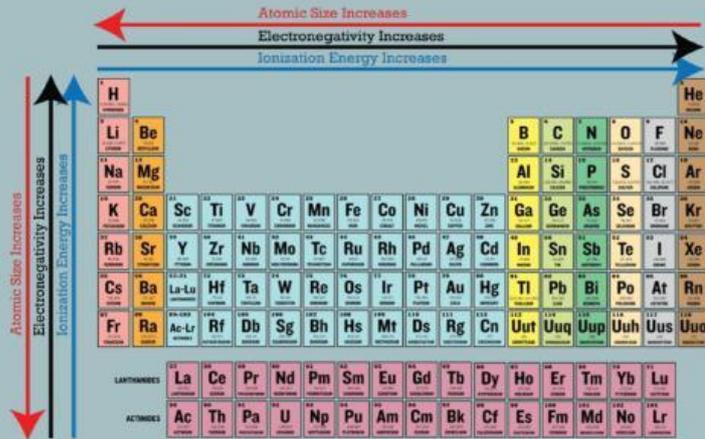
**Tuesday**

# Wednesday

Which element is the MOST electronegative?

The octet rule explains the stability of most covalently bonded molecules in terms of:

- A. ? noble-gas configurations
- B. ? ionization energies
- C. ? electronegativity differences
- D. ? electron affinities



# Thursday

In drawing Lewis structures, a single line (single bond) between two elements represents:

- A.  a shared electron
- B.  an octet of electrons
- C.  an unshared pair of electrons
- D.  a shared pair of electrons

Which of the following is an acceptable Lewis Structure for the diatomic nitrogen molecule?

- A.   $\ddot{\text{N}}=\ddot{\text{N}}$
- B.   $:\text{N}\equiv\text{N}:$
- C.   $\cdot\text{N}\equiv\text{N}\cdot$
- D.   $\ddot{\text{N}}\equiv\ddot{\text{N}}$

Which of the diatomic elements has a double bond between its atoms?

- A.  fluorine
- B.  nitrogen
- C.  hydrogen
- D.  oxygen

# Friday

The seven elements that occur as diatomic elements are:

- A.   $\text{Fe}_2, \text{Rn}_2, \text{O}_2, \text{He}_2, \text{Ne}_2, \text{C}_2, \text{Br}_2$
- B.   $\text{H}_2, \text{N}_2, \text{O}_2, \text{He}_2, \text{Ne}_2, \text{Cl}_2, \text{Br}_2$
- C.   $\text{H}_2, \text{N}_2, \text{O}_2, \text{He}_2, \text{Ne}_2, \text{C}_2, \text{Na}_2$
- D.   $\text{H}_2, \text{N}_2, \text{O}_2, \text{F}_2, \text{Cl}_2, \text{Br}_2, \text{I}_2$

In the correct Lewis structure for water, how many unshared pairs of electrons will oxygen have?

- A.  1
- B.  2
- C.  3
- D.  4