



# Weekly Warm Ups

Chemistry

# Monday

None

# Tuesday

None

# Wednesday

1. What is a binary ionic compound?

- A binary ionic compound is a compound that contains \_\_\_\_\_ different types of ions.

2. Identify the ions:

- The two types of ions found in a binary ionic compound are \_\_\_\_\_ ions and \_\_\_\_\_ ions. An example of a \_\_\_\_\_ ion is \_\_\_\_\_, and an example of a \_\_\_\_\_ ion is \_\_\_\_\_.

3. Practice Writing:

- The formula for the binary ionic compound formed between sodium (Na) and chlorine (Cl) is \_\_\_\_\_.

## Thursday

1. What would be the proper chemical formula for combining  $\text{Al}^{3+}$  and  $\text{Cl}^-$  :

1.  $\text{AlCl}_3$
2.  $\text{Al}_3\text{Cl}$
3.  $\text{AlCl}$
4.  $\text{Al}_3\text{Cl}_3$

2. When ions having a positive charge form bonds with ions having a negative charge, the charge on the resulting compound is:

# Friday

Write the formulas from the following graphic

Name		Ions
cesium nitride	→	$\text{Cs}^+ \begin{array}{c} \leftarrow \\ \rightarrow \end{array} \text{N}^{3-}$
barium oxide	→	$\text{Ba}^{2+} \begin{array}{c} \leftarrow \\ \rightarrow \end{array} \text{O}^{2-}$
calcium phosphide	→	$\text{Ca}^{2+} \begin{array}{c} \leftarrow \\ \rightarrow \end{array} \text{P}^{3-}$
iron(III) phosphide	→	$\text{Fe}^{3+} \begin{array}{c} \leftarrow \\ \rightarrow \end{array} \text{P}^{3-}$
calcium fluoride	→	$\text{Ca}^{2+} \begin{array}{c} \leftarrow \\ \rightarrow \end{array} \text{F}^-$
tin(IV) oxide	→	$\text{Sn}^{4+} \begin{array}{c} \leftarrow \\ \rightarrow \end{array} \text{O}^{2-}$