

DO NOW-MINI LAB #2

You have a bag and a small container on your desk. The bag has Calcium Chloride [CaCl_2] in it, the container has Baking Soda [NaHCO_3]. Pour the Baking Soda in the bag and crush and rub the two substances together. The reaction equation is:



Na

H

C

O

Ca

Cl

1. What are the reactants? (Select the FORMULAS)

$\text{NaHCO}_{3(s)}$ $\text{CaCl}_{2(s)}$ $\text{CaCO}_{3(s)}$ $\text{NaCl}_{(s)}$ $\text{H}_2\text{O}_{(l)}$ $\text{CO}_{2(g)}$

2. What are the products? (Select the FORMULAS)

$\text{NaHCO}_{3(s)}$ $\text{CaCl}_{2(s)}$ $\text{CaCO}_{3(s)}$ $\text{NaCl}_{(s)}$ $\text{H}_2\text{O}_{(l)}$ $\text{CO}_{2(g)}$

3. What liquid is produced? (Write the formula)

4. What gas is produced? (Write the formula)

5. Is the reaction balanced? YES NO

6. If the equation is not balanced, what's wrong and how could you fix it? Select where you would add coefficients

$\text{NaHCO}_{3(s)}$ $\text{CaCl}_{2(s)}$ $\text{CaCO}_{3(s)}$ $\text{NaCl}_{(s)}$ $\text{H}_2\text{O}_{(l)}$ $\text{CO}_{2(g)}$

7. Is the reaction (exothermic) letting heat out, or (endothermic) absorbing heat (getting cold)?

8. How can you tell this was a chemical change and not a physical change?

You can tell because there was a _____ produced.