1+		
1		
1	2+	Oxidation Number
1 H Hydrogen 1.008	2	Valence Electrons
	2	Family
3 Li Lithium 6.941	4 Be Beryllium 9.0122	
11 Na Sodium 22.990	12 Mg Magnesium 24.305	
19 K Potassium 39.098	20 Ca Calcium 40.078	

Some Polyatomic Ions and their Oxidation Numbers

1+	1-	2-	3-
ammonium (NH_4)	acetate ($C_2H_3O_2$)	carbonate (CO_3)	phosphate (PO_4)
	chlorate (ClO_3)	sulfate (SO_4)	
	hydroxide (OH)		
	nitrate (NO_3)		
	bicarbonate(HNO_3)		

						0
						8
Oxidation Number	3+	4+/-	3-	2-	1-	8
Valence Electrons	3	4	5	6	7	2 He Helium 4.003
Family	3	4	5	6	7	
	5 B Boron 10.81	6 C Carbon 12.011	7 N Nitrogen 14.007	8 O Oxygen 15.999	9 F Fluorine 18.998	10 Ne Neon 20.180
	13 Al Aluminum 26.98	14 Si Silicon 28.086	15 P Phosphorus 30.974	16 S Sulfur 32.06	17 Cl Chlorine 35.453	18 Ar Argon 39.948
	31 Ga Gallium 69.72	32 Ge Germanium 72.61	33 As Arsenic 74.922	34 Se Selenium 78.96	35 Br Bromine 79.904	36 Kr Krypton 83.80

8. The oxidation number of an ion is the number of electrons an atom will lose to become a **negative positive** ion or gain to become a **negative positive** ion and have a full outer energy level of electrons.

Combine the elements or polyatomic compounds and write the formula and the name

	Symbols & Oxidation #	Formula	Name of Compound
9.Sodium and bicarbonate			
10. Lithium and nitrogen			
11. Beryllium and hydroxide			
12. Magnesium and chlorine			
13. Calcium and nitrate			
14. Potassium and acetate			
15. Sodium and phosphate			
16. Beryllium and bromine			
17. Aluminum and sulfate			
18. Sodium and carbonate			

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