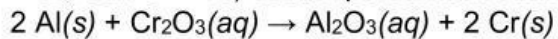


Sebanyak 8,1 gram logam aluminium direaksikan dengan 38 gram Cr_2O_3 ($A_r \text{ Al} = 27$, $\text{Cr} = 52$, dan $\text{O} = 16$) sesuai persamaan reaksi:



Berapa gram logam kromium yang dihasilkan?

Jawab

Menghitung mol Al

$$\text{mol Al} = \text{massa Al} : A_r \text{ Al}$$

$$= \quad :$$

$$= \quad \text{mol}$$

menghitung mol Cr_2O_3

$$M_r \text{Cr}_2\text{O}_3 = \quad \times A_r \text{Cr} + \quad \times A_r \text{O}$$

$$= \quad \times \quad + \quad \times \quad$$

$$=$$

$$\text{mol Cr}_2\text{O}_3 = \text{massa Cr}_2\text{O}_3 : M_r \text{Cr}_2\text{O}_3$$

$$= \quad :$$

$$= \quad \text{mol}$$

Reaksi	:	2Al(s)	+	$\text{Cr}_2\text{O}_3(\text{aq})$	$\rightarrow$	$\text{Al}_2\text{O}_3(\text{aq})$	+	2Cr(s)
mula-mula (mol)	:							
bereaksi (mol)	:							
akhir (mol)	:							

$$\text{mol Cr yang dihasilkan} = \quad \text{mol}$$

$$\text{massa Cr yang dihasilkan} = \text{mol Cr} \times M_r \text{Cr}$$

$$= \quad \times$$

$$= \quad \text{gram}$$