Multiple Choice: Choose the letter of the correct answer.

_____ **1.** Which term indicates how strongly an atom attracts the electrons in a chemical bond?
A) alkalinity C) electronegativity
B) atomic mass. D) activation energy

_____ **2.** An ionic compound is formed when there is a reaction between the elements
A) strontium and chlorine C) nitrogen and oxygen
B) hydrogen and chlorine D) sulfur and oxygen

_____ **3.** Which element has an atom with the greatest tendency to attract electrons in a chemical bond?
A) carbon C) silicon
B) chlorine D) sulfur

_____ **4.** How many electrons are needed in the outer energy levels of an atom to be stable?
A. 2 B. 6 C. 4 D. 8

_____ **5.** A charged particle that has gained or lost electrons is called a _____.
A. molecule B. Isotope C. Ion D. Element

_____ **6.** Nitrogen will _____valence electrons when forming an ionic bond.
A. gain 1 B. gain 3 C. lose 1 D. lose 3

_____ **7.** Ionic bonds form between metals and _____.
A. metals B. non-metals C. metalloids D. none

_____ **8.** If I am in group 17 how many valence electrons do I have?
A. 17 B. 7 C. 1 D. none

_____ **9.** When an atom loses a valence electron, it becomes a(n) _____ion.
A. positive B. neutral C. negative D. polyatomic

_____ **10.** In order to have a stable arrangement of 8 valence electrons, metal atoms are likely to _____electrons.
A. no charge B. negative C. positive D. it depends

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