

ENTROPY

State whether the entropy of the system

INCREASES (I) (S gets more positive); DECREASES (D) (S gets more negative); REMAINS NEAR CONSTANT (RC) (S stays around 0)

a	a solid crystallising from solution	I	D	RC
b	a solid melting	I	D	RC
c	a solid subliming	I	D	RC
d	two liquids mixing together but not reacting	I	D	RC
e	$\text{CaCO}_3(\text{s}) \rightarrow \text{CaO}(\text{s}) + \text{CO}_2(\text{g})$	I	D	RC
f	water electrolysing to O_2 and H_2	I	D	RC
g	$\text{O}_2(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$	I	D	RC
h	$\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$	I	D	RC
i	$\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{NO}(\text{g})$	I	D	RC