

Ionic Nomenclature Practice 3 (Polyatomics & Transition Metals)

Roman Numeral Needed?	Name	Ions	Formula
NO	sodium perchlorate	Na^+ ClO_4^-	NaClO_4
		Cr^{+2} S^{-2}	
			Li_3P
			MnN
	aluminum bromide		
			Sn_2O_3
	zinc oxide		
		Fe^{+3} I^-	
	lead(II) sulfate		
			RbI
		Al^{+3} Se^{-2}	
	beryllium fluoride		
			TiCl_2
		Sr^{+2} P^{-3}	
	lead(IV) acetate		

			KBrO ₃
	cobalt(II) iodide		
			Ag ₂ Se
		K ⁺ S ⁻²	
	iron(III) phosphate		
			Mg(IO ₃) ₂
		Ca ⁺² N ⁻³	
			Cs ₂ O
	gallium hydride		

Metals Review:

Metals are located on the _____ of the periodic table.

Metals have _____ electronegativities which means they are _____ at taking other atoms' electrons.

Metals have _____ ionization energies which means it is _____ to take their electrons.

Metals tend to _____ electrons to become _____ charged _____

Nonmetals Review:

Nonmetals are located on the _____ of the periodic table.

Nonmetals have _____ electronegativities which means they are _____ at taking other atoms' electrons.

Nonmetals have _____ ionization energies which means it is _____ to take their electrons.

Nonmetals tend to _____ electrons to become _____ charged _____.