

**Ionic Nomenclature Practice 3 (Polyatomics & Transition Metals)**

Roman Numeral Needed?	Name	Ions	Formula
NO	sodium perchlorate	$\text{Na}^+$ $\text{ClO}_4^-$	$\text{NaClO}_4$
		$\text{Cr}^{+2}$ $\text{S}^{-2}$	
			$\text{Li}_3\text{P}$
			$\text{MnN}$
	aluminum bromide		
			$\text{Sn}_2\text{O}_3$
	zinc oxide		
		$\text{Fe}^{+3}$ $\text{I}^-$	
	lead(II) sulfate		
			$\text{RbI}$
		$\text{Al}^{+3}$ $\text{Se}^{-2}$	
	beryllium fluoride		
			$\text{TiCl}_2$
		$\text{Sr}^{+2}$ $\text{P}^{-3}$	
	lead(IV) acetate		

			KBrO <sub>3</sub>
	cobalt(II) iodide		
			Ag <sub>2</sub> Se
		K <sup>+</sup> S <sup>-2</sup>	
	iron(III) phosphate		
			Mg(IO <sub>3</sub> ) <sub>2</sub>
		Ca <sup>+2</sup> N <sup>-3</sup>	
			Cs <sub>2</sub> O
	gallium hydride		

**Metals Review:**

Metals are located on the \_\_\_\_\_ of the periodic table.

Metals have \_\_\_\_\_ electronegativities which means they are \_\_\_\_\_ at taking other atoms' electrons.

Metals have \_\_\_\_\_ ionization energies which means it is \_\_\_\_\_ to take their electrons.

Metals tend to \_\_\_\_\_ electrons to become \_\_\_\_\_ charged \_\_\_\_\_

**Nonmetals Review:**

Nonmetals are located on the \_\_\_\_\_ of the periodic table.

Nonmetals have \_\_\_\_\_ electronegativities which means they are \_\_\_\_\_ at taking other atoms' electrons.

Nonmetals have \_\_\_\_\_ ionization energies which means it is \_\_\_\_\_ to take their electrons.

Nonmetals tend to \_\_\_\_\_ electrons to become \_\_\_\_\_ charged \_\_\_\_\_.