

STRONG BASE

Given the concentration of $\text{NaOH} = 0.2 \text{ M}$.
Calculate the pH of of the base solution

$$\begin{aligned}\text{pOH} &= -\log [\text{OH}^-] \\ &= -\log \underline{\hspace{2cm}} \\ &= \underline{\hspace{2cm}}\end{aligned}$$

$$\begin{aligned}\text{pH} &= 14 - \text{pOH} \\ &= 14 - \underline{\hspace{2cm}} \\ &= \underline{\hspace{2cm}}\end{aligned}$$

STRONG BASE

Given the concentration of $\text{Ba}(\text{OH})_2 = 0.2 \text{ M}$.
Calculate the pH of the base solution

$$\begin{aligned}\text{pOH} &= -\log [\text{OH}^-] \\ &= -\log \underline{\hspace{2cm}} \\ &= \underline{\hspace{2cm}}\end{aligned}$$

$$\begin{aligned}\text{pH} &= 14 - \text{pOH} \\ &= 14 - \underline{\hspace{2cm}} \\ &= \underline{\hspace{2cm}}\end{aligned}$$