Elite Private School

Worksheet

ACID AND BASE DISSOCIATION.

Using the correct information from each table, complete the chemical equation and dissociation expressions for each **ACID**.

A. $H_3PO_4(aq) + \quad (l) \rightleftharpoons \quad (aq) + \quad (aq)$

$K_a =$ _____

H_3O^+	H_2O	$H_2PO_4^-$
$[H_3O^+]$	$[H_3PO_4]$	$[H_2PO_4^-]$

B. $H_2Te(aq) + \quad (l) \rightleftharpoons \quad (aq) + \quad (aq)$

$K_a =$ _____

H_3O^+	H_2O	HTe^-
$[H_3O^+]$	$[H_2Te]$	$[HTe^-]$

C. $H_4P_2O_7(aq) + \quad (l) \rightleftharpoons \quad (aq) + \quad (aq)$

$K_a =$ _____

H_3O^+	H_2O	$H_3P_2O_7^-$
$[H_3O^+]$	$[H_4P_2O_7]$	$[H_3P_2O_7^-]$

D. $HC_3H_5O_3(aq) + \quad (l) \rightleftharpoons \quad (aq) + \quad (aq)$

$K_a =$ _____

H_3O^+	H_2O	$C_3H_5O_3^-$
$[H_3O^+]$	$[HC_3H_5O_3]$	$[C_3H_5O_3^-]$

E. Write the chemical equation and dissociation expression for the weak acid, H_2SO_3.

LIVEWORKSHEETS

F. Write the following:

a. **The equation for the acid H_2CO_3;**
b. **The equation for the conjugate base of the acid**
c. **The K_a expression for the acid**
d. **The K_b expression for the conjugate base.**

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