

Identification of Limiting Reactant

Q#1: The reaction between NaOH and H₂SO₄ produces Na₂SO₄ and H₂O as products. If 4 grams of NaOH reacts with 9.8 grams of H₂SO₄ then find the limiting reactant.

2NaOH + H ₂ SO ₄ → Na ₂ SO ₄ + 2H ₂ O		
1. Moles of reactants	$NaOH = \frac{4}{40} =$	$H_2SO_4 = \frac{9.8}{98} =$
2. Divide moles on coefficient of reactant.	$NaOH = \frac{1}{2}$	$H_2SO_4 = \frac{1}{1}$
3. Limiting Reactant		

Q#2: The reaction between Hydrogen and Oxygen produces water.

- Balance the Chemical Equation for the reaction.
- If 0.8 grams of H₂ reacts with 1.6 grams of O₂ then find the limiting reactant.

H ₂ + O ₂ → H ₂ O		
1. Moles of reactants	$Hydrogen = \frac{0.8}{2} =$	$Oxygen = \frac{1.6}{32} =$
2. Divide moles on coefficient of reactant.	$Hydrogen = -$	$Oxygen = - =$
3. Limiting Reactant		