

Identification of Limiting Reactant

Q#1: The reaction between NaOH and H₂SO₄ produces Na₂SO₄ and H₂O as products. If 4 grams of NaOH reacts with 9.8 grams of H₂SO₄ then find the limiting reactant.

2NaOH + H ₂ SO ₄ → Na ₂ SO ₄ + 2H ₂ O		
1. Moles of reactants	NaOH = $\frac{40}{40} =$	H ₂ SO ₄ = $\frac{98}{98} =$
2. Divide moles on coefficient of reactant.	NaOH = $\frac{1}{2}$	H ₂ SO ₄ = $\frac{1}{1}$
3. Limiting Reactant		

Q#2: The reaction between Hydrogen and Oxygen produces water.

- Balance the Chemical Equation for the reaction.
- If 0.8 grams of H₂ reacts with 1.6 grams of O₂ then find the limiting reactant.

H ₂ + O ₂ → H ₂ O		
1. Moles of reactants	Hydrogen = $\frac{2}{2} =$	Oxygen = $\frac{32}{32} =$
2. Divide moles on coefficient of reactant.	Hydrogen = $\frac{1}{2} =$	Oxygen = $\frac{1}{1} =$
3. Limiting Reactant		