

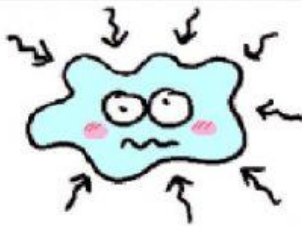
Latent Heat of Vaporization

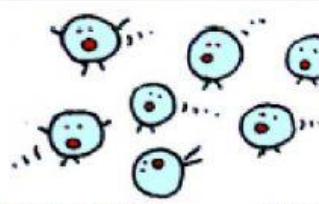


Oh no!
not again!

By Esther

What will happen if we continue heating up the water?

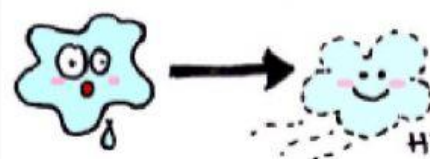
1. 
At 0°C , the water absorbs heat...

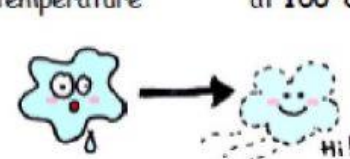
2. 
Heat energy $\rightarrow$ energy
The water molecules gain kinetic energy and move around

3. 
The temperature of the water rises from 0°C to the point of water 100°C .

4. At 100°C , the water starts

am I disappearing again??

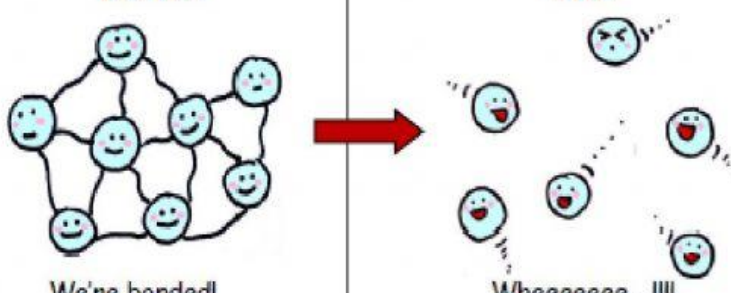
5. It does not disappear!

The water changes into steam! This process is called vaporization.

6. During phase change, the temperature stays at 100°C .

There's no change in temperature!

Q: If there is no change in temperature of water and steam during phase change, what happen to the heat absorbed by water?

Tick the correct statement:

- ☐ Weaken the bonds between the water molecules.
- ☐ Increase the kinetic energy of water molecules

7. 

The heat required to change 1 kg of substance from liquid to steam is

Latent heat of Vaporization, ℓ (J/kg)