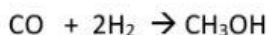


NAME: _____

STOICHIOMETRY WORKSHEET: Percentage Yield

Solve the following stoichiometry problems and write the correct answer with units.
Round the number to 1 decimal point and remember you need to have balanced equations to work with.

1. Calculate the indicated quantity (X) for each of the following:
 - a. Theoretical yield= 20 g Actual yield= 15 g Percentage yield= X
ANSWER: _____
 - b. Theoretical yield= 1.5 g Actual yield= X g Percentage yield= 90%
ANSWER: _____
 - c. Theoretical yield= 5 g Actual yield= 4.75 g Percentage yield= X
ANSWER: _____
2. Methanol can be produced through the reaction of CO and H₂ in the presence of a catalyst.



If 75 g of CO react to produce 68.4 g CH₃OH, what is the percentage yield of CH₃OH?
ANSWER: _____

3. Aluminum reacts with excess Copper (II) Sulfate according to the reaction given below. If 1.85 g of Al react, and the percentage yield of Cu is 56.6%, what mass of Cu is produced?



ANSWER: _____

4. Quicklime, CaO, can be prepared by roasting limestone, CaCO₃, according to the following reaction:



When 200 g of CaCO₃ are heated, the actual yield of CaO is 105 g. What is the percentage yield?

ANSWER: _____