Introduction to Kinetic Molecular Theory

Question 1

Kinetic Molecular Theory states that gas particles are in constant motion and exhibit bounced elastic collisions.
The average **kinetic energy** of all ion of gas particles is **directly proportional** to absolute **temperature** only.

Instructions: Use the above statesmen to complete the paragraph below

Kinetic _____ Theory states that _____ _____ are in _____ _____ and exhibit _____ elastic _____.

The average _____ **energy** of _____ _____ gas _____ is **directly** _____ ___ absolute _____ only.

Question 2

Label and Identify Solid, Liquid or Gas
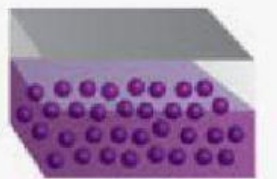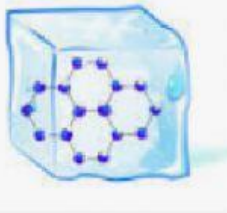
What is it? Vibrate on the spot, cannot move from one place to another	**What is it?** Move and slide around each other	**What is it?** Move very quickly in all directions

Question 3

Instructions: Complete by using the word "more" or the word "less"

- The faster the molecules move the _____ Temperature measurement.

- The slower the molecules move the _____ Temperature measurement.

Question 4

Instructions: Complete by using the word "more" or the word "less"

- The less average kinetic energy the _____ temperature measurement

- The more average kinetic energy the _____ temperature measurement

Question 5

The kinetic molecular theory

1. All matter is made of particles

2. Particles are in constant , random motion.

3. Particles constantly collide with each other and with the walls of their container

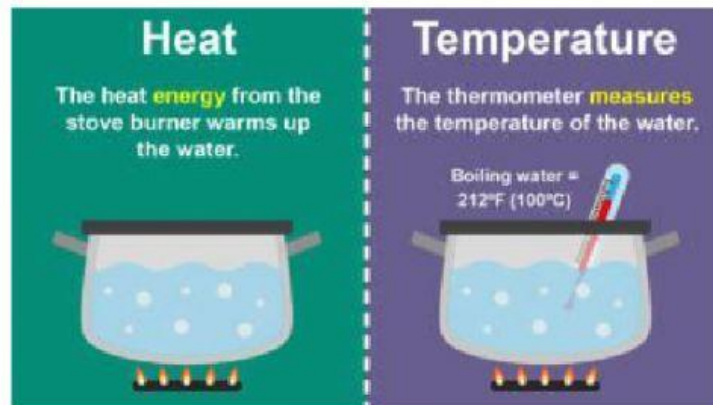**Instructions**: Read and Copy statements 1, 2 and 3.

1. _____

2. _____

3. _____

_____

4. Are you made of particles? Yes or No _____

Question 6

Heat vs Temperature

Heat
The heat energy from the stove burner warms up the water.

Temperature
The thermometer measures the temperature of the water.

Boiling water = 212°F (100ºC)

Instructions: Use the above picture to answer thje questions below

1. What is Heat? _____
_____

2. What is Temperature? _____
_____

3. Is Heat and Temperature the same? Yes or No _____