Writing Chemical Formulas and Names: Remediation Worksheet

1. When it comes to naming, compounds that are made up of a metal and nonmetal are considered _____.
2. When it comes to naming, compounds that are made up of a nonmetal and nonmetal are considered _____.
3. Nonmetals are located _____ on the periodic table.
4. Metals are located _____ on the periodic table.
5. How many valence electrons do elements in group 2 have? _____
6. How many valence electrons do elements in group 16 have? ____
7. Choose the type of compound that applies to the rules below for writing chemical formulas.
   a. Prefixes are used to determine the subscripts (or number of each atom) in the chemical formula. _____
   b. Charges are "Crisscrossed" to determine the subscripts. If necessary, they are put in the simplest form. _____

8. Complete the table below:

Name	Ionic or covalent	Formula (write a 1 if there is only 1 even though we don't usually write 1 in formulas)
Diphosphorus pentoxide		P O
Aluminum chloride		Al Cl
Nitrogen trioxide		N O
Aluminum nitrate		Al (NO_3)
Potassium chloride		K Cl
Calcium oxide		Ca O
Fluorine heptabromide		F Br
Calcium phosphide		Ca P
Carbon tetrafluoride		C F

9. Choose the type of chemical compound that matches the rule for writing names.
    a. The first element's name is written as is and the second compound's name is changed to end in "ide". No prefixes are used. _____
    b. T The first element's name is written as is and the second compound's name is changed to end in "ide". Prefixes are used to indicate the number of each element.

    _____

10. Complete the table below:

Formula	**Ionic or covalent**	**Name (be careful of spelling; no caps)**
BaF_2		
N_2H_4		
$Ca_3(PO_4)_2$		
MgH_2		
CO_2		
NH_3		
S_2O_8		
Li_2O		
$Mg(OH)_2$		