

Carbon & Its Compounds

Class-VII

WS-4

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Drag the correct option to the correct blank:

Covalent	2-Dimentional
Diamond	Carbon
Three	Graphite
Van der waal's force	Silicon carbide
3-Dimentional	Four

This exercise is based on properties of diamond & graphite:

- _____ allotrope of carbon is one of the hardest naturally occurring substance.
- _____ allotrope of carbon is naturally found as plumbago.
- Diamond has a _____ structure.
- In Graphite each carbon is bonded to _____ other carbon by covalent bond.
- Graphite has a _____ structure.
- In Diamond each carbon is bonded to _____ other carbon by covalent bond.
- Graphite can be artificially prepared by heating _____
- Diamond can be artificially prepared by heating _____ under high temperature & pressure
- Graphite is soft because of flat layers joined together by _____
- Diamond is hard because of carbon atom joined together by _____