# Equation writing and stoichiometry:

a) Balance the following equation and name each of the chemicals in the box below its formula. For a coefficient of one, leave the box blank.

$Pb(NO_3)_{2\ (aq)}$ +        $KI_{(aq)}$              →              $KNO_{3\ (aq)}$      +              $PbI_{2\ (s)}$

b) If 0.00050 kg of $Pb(NO_3)_{2\ (aq)}$ reacted, how many grams of $PbI_{2\ (s)}$ would be produced?

Molar mass of $Pb(NO_3)_2$ to the nearest tenth:              g/mol

Moles of $Pb(NO_3)_{2\ (aq)}$ [3 significant figures]:              mol

Moles of $PbI_{2\ (s)}$ [3 significant figures]:              mol

Grams of $PbI_{2\ (s)}$ [correct number of significant figures for final answer]: