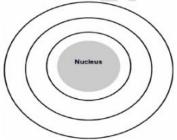


Atomic Structure 101 Video Review

Follow along with the	video and complete the Bohr Model for Hydrogen below:
	Protons =
	Neutrons =
Nucleus	Electrons =
	Atomic # =
	Mass # =
	Valence electrons =
. How do you find the	nass number for an element?
	nass number for an element? video and complete the Bohr Model for Sodium below:
	video and complete the Bohr Model for Sodium below:
	video and complete the Bohr Model for Sodium below: Protons =
	video and complete the Bohr Model for Sodium below: Protons = Neutrons =
	video and complete the Bohr Model for Sodium below: Protons = Neutrons = Electrons =
	video and complete the Bohr Model for Sodium below: Protons = Neutrons = Electrons = Atomic # =
	video and complete the Bohr Model for Sodium below: Protons = Neutrons = Electrons = Atomic # = Mass # =
Follow along with the	video and complete the Bohr Model for Sodium below: Protons = Neutrons = Electrons = Atomic # = Mass # =



Protons = ______

Neutrons = _____

Electrons = ____

Atomic # = ____

Mass # = ____

Valence electrons = ____

7. Follow along with the video and complete the Bohr Model for Helium below:

	Protons =
	Neutrons =
Nucleus	Electrons =
	Atomic # =
	Mass # =
	Valence electrons =

8. Follow along with the video and complete the Bohr Model for Calcium below:

	Protons =
	Neutrons =
Nucleus	Electrons =
	Atomic # =
	Mass # =
	Valence electrons =

Atomic Structure Quiz

- 1. What 3 subatomic particles are located in an atom? __ 2. What 2 subatomic particles are located in the nucleus of an atom? 3. What subatomic particle has a positive charge? ______ Negative charge? _____
- Neutral/no charge? Electrons are located outside of the nucleus in the ______
- 5. You calculate the mass number by adding _
- 6. How do you find the number of valence electrons?
- 7. The number of protons is the same as the _ What subatomic particle is the identify of an element?
- 9. Why is the nucleus of an atom positive? _
- 10. What 2 subatomic particles cancel each other out? ____