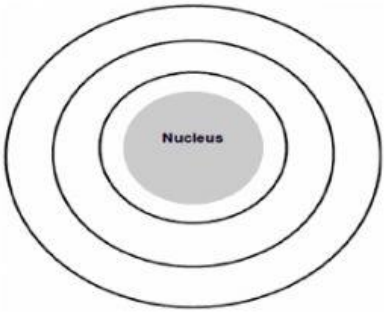
Atomic Structure 101 Video Review

1. Who created the Bohr Model? _____
2. What is the Bohr Model? _____

   _____
3. Follow along with the video and complete the Bohr Model for Hydrogen below:



Protons = _____

Neutrons = _____

Electrons = _____

Atomic # = _____

Mass # = _____

Valence electrons = ____

4. How do you find the mass number for an element? _____
5. Follow along with the video and complete the Bohr Model for Sodium below:

Protons = _____

Neutrons = _____

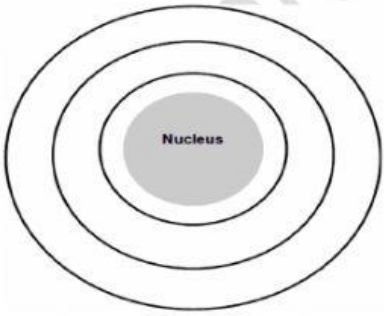
Electrons = _____

Atomic # = _____

Mass # = _____

Valence electrons = ____

6. Follow along with the video and complete the Bohr Model for Neon below:



Protons = _____

Neutrons = _____

Electrons = _____

Atomic # = _____

Mass # = _____

Valence electrons = ____

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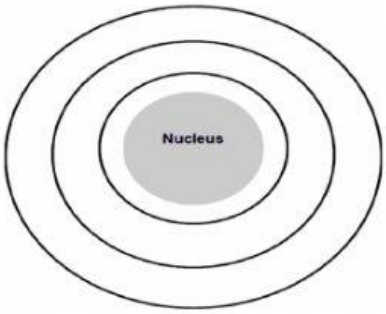
7. Follow along with the video and complete the Bohr Model for Helium below:



Protons = _____

Neutrons = _____

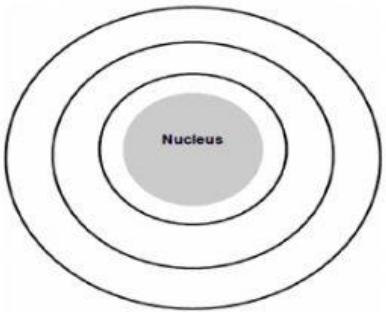
Electrons = _____

Atomic # = _____

Mass # = _____

Valence electrons = ___

8. Follow along with the video and complete the Bohr Model for Calcium below:



Protons = _____

Neutrons = _____

Electrons = _____

Atomic # = _____

Mass # = _____

Valence electrons = ___

Atomic Structure Quiz

1. What 3 subatomic particles are located in an atom? _____
2. What 2 subatomic particles are located in the nucleus of an atom? _____
3. What subatomic particle has a positive charge? _____ Negative charge? _____ Neutral/no charge? _____
4. Electrons are located outside of the nucleus in the _____
5. You calculate the mass number by adding _____
6. How do you find the number of valence electrons? _____
7. The number of protons is the same as the _____
8. What subatomic particle is the identify of an element? _____
9. Why is the nucleus of an atom positive? _____
10. What 2 subatomic particles cancel each other out? _____

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