

Name:  
Stoichiometry



a) How many **moles** of iron would be needed to react with 3.82 **moles** of oxygen?

$\underline{\hspace{2cm}} = \text{mol Fe}$

b) What **mass** of iron (III) oxide can be produced from 1.35 **moles** Fe?

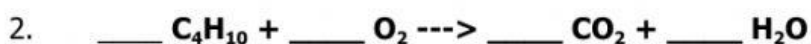
$\underline{\hspace{2cm}} = \text{g Fe}_2\text{O}_3$

c) How many **moles** of  $\text{O}_2$  are needed to produce 347 **g** of  $\text{Fe}_2\text{O}_3$ ?

$\underline{\hspace{2cm}} = \text{mol O}_2$

d) What **mass** of iron (III) oxide can be produced from 135 **g** Fe?

$\underline{\hspace{2cm}} = \text{g Fe}_2\text{O}_3$



a) When 6.24 **moles** of  $\text{O}_2$  are reacted, how many **moles** of carbon dioxide are produced?

$\underline{\hspace{2cm}} = \text{mol CO}_2$

b) How many **grams** of  $\text{C}_4\text{H}_{10}$  would produce 88 **grams** of water?

$\underline{\hspace{2cm}} = \text{g C}_4\text{H}_{10}$

6.24 mol $\text{O}_2$	88 g $\text{H}_2\text{O}$
10 $\text{H}_2\text{O}$	58.14 g
13 $\text{O}_2$	44.01 g
2 $\text{C}_4\text{H}_{10}$	32.00 g
8 $\text{CO}_2$	18.02 g
1 mole	6.022E23
1 mole	