

Name:  
Stoichiometry



a) How many **moles** of iron would be needed to react with 3.82 **moles** of oxygen?



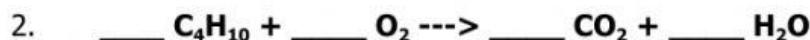
b) What **mass** of iron (III) oxide can be produced from 1.35 **moles** Fe?



c) How many **moles** of  $\text{O}_2$  are needed to produce 347 **g** of  $\text{Fe}_2\text{O}_3$ ?



d) What **mass** of iron (III) oxide can be produced from 135 **g** Fe?



a) When 6.24 **moles** of  $\text{O}_2$  are reacted, how many **moles** of carbon dioxide are produced?

6.24 mol  $\text{O}_2$

88 g  $\text{H}_2\text{O}$

10  $\text{H}_2\text{O}$

58.14 g

13  $\text{O}_2$

44.01 g

2  $\text{C}_4\text{H}_{10}$

32.00 g

8  $\text{CO}_2$

18.02 g

1 mole

6.022E23

1 mole

b) How many **grams** of  $\text{C}_4\text{H}_{10}$  would produce 88 **grams** of water?

