

Name: _____ Date: _____ Number: _____

Mole Ratio / Mole-Mole Stoichiometry Worksheet

1. Balance the equation and answer the following: $\text{N}_2 + \text{H}_2 \rightarrow \text{NH}_3$

a. If you used 1 mole of N_2 , how many moles of NH_3 could be produced?

_____ =

b. If 10 moles of NH_3 were produced, how many moles of N_2 would be required?

_____ =

c. If 3.00 moles of H_2 were used, how many moles of NH_3 would be made?

_____ =

d. If 0.600 moles of NH_3 were produced, how many moles of H_2 are required?

_____ =

2. For the reaction: $\text{P} + \text{S} \rightarrow \text{P}_2\text{S}_5$

a. How many moles of phosphorus are needed to react with 0.125 moles of sulfur?

_____ =

b. How many moles of phosphorous pentasulfide are produced from 0.081 mol of sulfur in the above reaction?

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