CLASS	10 SCIENCE E/M	UNIT NO	10
TYPE OF QUESTION	CHOOSE THE CORRECT ANSWER.	MARK	10

1. $H_2(g) + Cl_2(g) \rightarrow 2HCl(g)$ is a

a. Decomposition Reaction b. Combination Reaction

c. Single Displacement Reaction d. Double Displacement Reaction

2. Photolysis is a decomposition reaction caused by _____

a. heat b. electricity c. light d. mechanical energy

3. A reaction between carbon and oxygen is represented by

$C(s) + O_2(g) \rightarrow CO_2(g)$ + Heat.

In which of the type(s), the above reaction can be classified?

(i) Combination Reaction (ii) Combustion Reaction

(iii) Decomposition Reaction (iv) Irreversible Reaction

a. i and ii b. i and iv c. i, ii and iii d. i, ii and iv

4. The chemical equation

$Na_2SO_4(aq) + BaCl_2(aq) \rightarrow BaSO_4(s)\downarrow + 2NaCl(aq)$

represents which of the following types of reaction?

a.Neutralisation b.Combustion

c. Precipitation d. Single displacement

5. Which of the following statements are correct about a chemical equilibrium?

(i) It is dynamic in nature (ii) The rate of the forward and backward reactions are equal at equilibrium (iii) Irreversible reactions do not attain chemical equilibrium

(iv) The concentration of reactants and products may be different

a. i, ii and iii b. i, ii and iv

c. ii, iii and iv d. i, iii and iv

6. A single displacement reaction is represented by

$X(s) + 2HCl(aq) \rightarrow XCl_2(aq) + H_2(g)$. Which of the following(s) could be X.

(i) Zn (ii) Ag (iii) Cu (iv) Mg.

Choose the best pair.

a. i and ii b. ii and iii c. iii and iv d. i and iv

7. Which of the following is not an "element + element → compound" type reaction?

a. $C(s) + O_2(g) \rightarrow CO_2(g)$ b. $2K(s) + Br_2(l) \rightarrow 2KBr(s)$

c. $2CO(g) + O_2(g) \rightarrow 2CO_2(g)$ d. $4Fe(s) + 3O_2(g) \rightarrow 2Fe_2O_3(s)$

8. Which of the following represents a precipitation reaction?

a. $A(s) + B(s) \rightarrow C(s) + D(s)$ b. $A(s) + B(aq) \rightarrow C(aq) + D(l)$

c. $A(aq) + B(aq) \rightarrow C(s) + D(aq)$ d. $A(aq) + B(s) \rightarrow C(aq) + D(l)$

9. The pH of a solution is 3. Its [OH–] concentration is

a. 1×10^{-3} M b. 3 M c. 1×10^{-11} M d. 11 M

10. Powdered CaCO3 reacts more rapidly than flaky $CaCO_3$ because of _____.

a. large surface area b. high pressure c. high concentration d. high temperature.

LIVEWORKSHEETS