



Calculating Percent Composition by Mass Video Notes

1. What is the percent composition by mass of Hydrogen in H_2O ?

percent composition by mass:
(percent by mass) $\frac{\text{total molar mass of just the element}}{\text{molar mass of the whole compound}} \times 100\%$

Calculate molar mass of H_2O - $H = 2 \times (\quad) + O = 1 \times (\quad) = \text{molar mass of } H_2O = \quad$

Write down the notes at the 3-minute mark of the video to solve the problem.

What is the percent composition by mass of Hydrogen in H_2O ? = $\quad$

2. Calculate the percent by mass of Iron in $FeCl_3$

Calculate the molar mass of $FeCl_3$ = $\quad$ Calculate the molar mass of Fe = $\quad$

Input your numbers in the following formula to solve.

$$\frac{\text{total molar mass of just the element}}{\text{molar mass of the whole compound}} \times 100\%$$

The percent by mass of Iron in $FeCl_3$ = $\quad$

Learning Target: I can calculate percent composition by mass.

3. Determine the percent by mass of Carbon in $C_6H_{10}O_4$.

Calculate the molar mass of $C_6H_{10}O_4$ = _____ Calculate the molar mass of Carbon = _____

Input your numbers in the following formula to solve.

$$\frac{\text{total molar mass of just the element}}{\text{molar mass of the whole compound}} \times 100\%$$

The percent by mass of Carbon in $C_6H_{10}O_4$ = _____