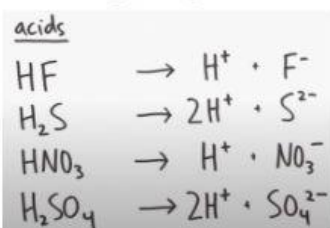


SC2. E. I can ask questions about chemical names to identify patterns in IUPAC nomenclature in order to predict chemical names for ionic, acidic, and inorganic covalent compounds.

Writing & Naming ACIDS Video Notes



1. An acid is a compound in which one or more _____ are bonded to a _____ ion.
When naming acids just like in any ionic



2. Groups of elements that together have a charge are called _____ ions. Just like in any ionic formula the positive charge and the negative charge have to _____ out.
3. The name of an acid is based on the name of the _____ ion that is part of the acid.
4. There are two types of acids. One that have _____ and ones that don't have _____

<u>negative ion</u>	<u>acid</u>
-ide	hydro _____ ic acid
chloride	hydrochloric acid

Naming Acids without Oxygen: Write down what is shown in the video for this part.

HCl = _____ acid

HBr = _____ acid

SC2. E. I can ask questions about chemical names to identify patterns in IUPAC nomenclature in order to predict chemical names for ionic, acidic, and inorganic covalent compounds.

<u>negative ion</u>	<u>acid</u>
-ate	-ic acid
nitrate	nitric acid

Naming Acids with Oxygen (Oxoacids): Write down what is shown in the video for this part.

HNO_3 = _____ acid

H_2CO_3 = _____ acid

<u>negative ion</u>	<u>acid</u>
-ite	-ous acid
nitrite	nitrous acid

HNO_2 = _____ acid

HCrO_2 = _____ acid

<u>Polyatomic Ions</u>	
Carbonate	CO_3^{2-}
Chromite	CrO_2^-
Hypochlorite	ClO^-
Nitrate	NO_3^-
Nitrite	NO_2^-
Permanganate	MnO_4^-
Phosphate	PO_4^{3-}
Phosphite	PO_3^{3-}
Sulfate	SO_4^{2-}
Sulfite	SO_3^{2-}

<u>negative ion</u>	<u>acid</u>	<u>example</u>
-ide	hydro_____ic acid	chloride hydrochloric acid
-ate	-ic acid	nitrate nitric acid
-ite	-ous acid	nitrite nitrous acid

My r_____ has _____ l_____s.

I _____ something _____ ky.

Spr_____ is delici_____.

Important Exception!

<u>acids</u>	<u>negative ions</u>
H_3PO_4	phosphate (PO_4^{3-})
H_3PO_3	phosphite (PO_3^{3-})
H_2SO_4	sulfate (SO_4^{2-})
H_2SO_3	sulfite (SO_3^{2-})

HMnO_4 = _____ acid

HClO = _____ acid

So how do we name acids?

Take a look at the _____ ion they have in them, we look at the _____ of that negative ion, and we use the _____ to figure out what we call the acid.

<u>Polyatomic Ions</u>	
Carbonate	CO_3^{2-}
Chromite	CrO_2^-
Hypochlorite	ClO^-
Nitrate	NO_3^-
Nitrite	NO_2^-
Permanganate	MnO_4^-
Phosphate	PO_4^{3-}
Phosphite	PO_3^{3-}
Sulfate	SO_4^{2-}
Sulfite	SO_3^{2-}

Created By: Chivas & Jordan Spivey