QUIZIZZ

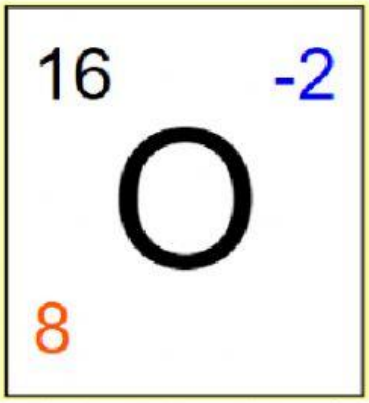
NAME :

CLASS :

DATE :

Unit 1: Atomic Structure
30 Questions

1.



What is Oxygen's atomic number?

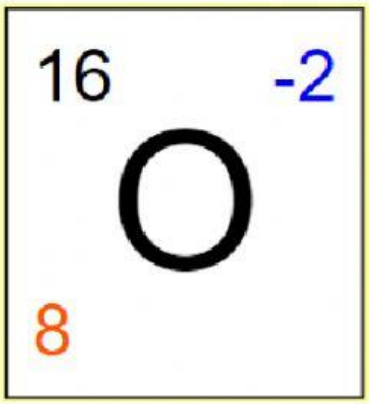
-2

8

10

16

2.



What is this atom's atomic mass?

8

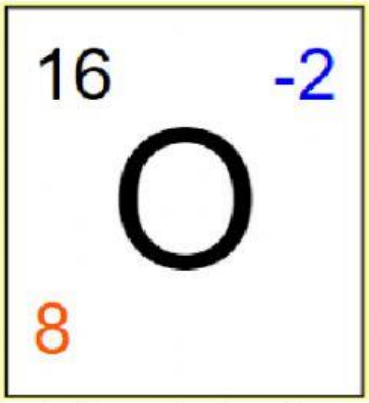
10

-2

16

LIVEWORKSHEETS

3.



What is this atom's charge?

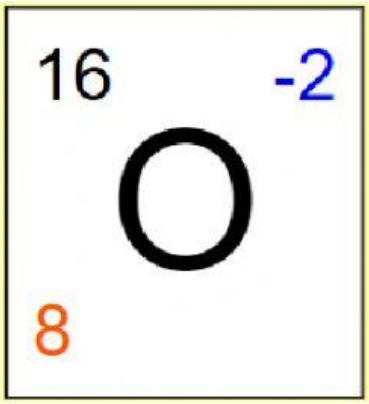
-2

8

10

16

4.



How many neutron's does this atom have?

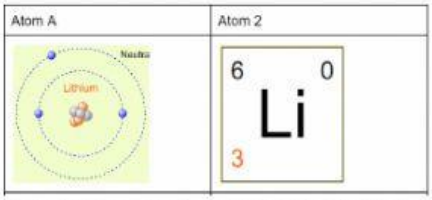
10

8

-2

16

5.



Determine the relationship between the following atoms:

The same atom

A different element

Isotopes of the same element

LIVEWORKSHEETS

6. What is the charge of electrons?

Electrons have no charge (0)

Electrons are positively (+) charged

Electrons are negatively (-) charged

7. What is the mass of a proton?

0

1

-1

8. Who discovered electrons?

Alexander the Great

J.J Thomson

Niels Bohr

9. What is the charge of a proton?

–

0

+

10. What is the charge of a neutron?

–

+

0

11. What is the mass of a neutron?

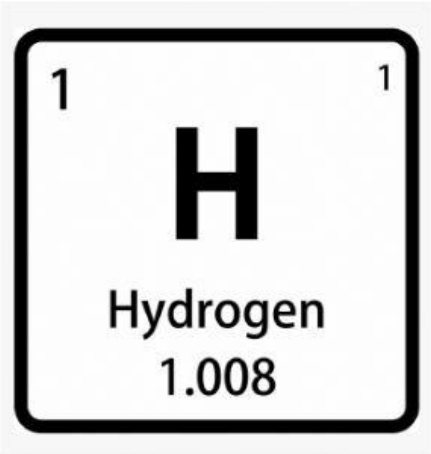
0

-1

1

LIVEWORKSHEETS

12.



How many NEUTRON are there in this Hydrogen?

0

2

1

13. What makes up the nucleus?

Electrons+ Protons

Protons + Neutrons

Neutrons + Electrons

14. What are three subparticles of an atom?

Proton, Neutron, Electron

One Oxygen and two Hydrogen. H_2O

Hydrogen, Oxygen, Helium

15. Where do electrons found?

Electrons are not moving

Around the nucleus

In the center of the atom

16. Where are protons found?

Around the nucleus

In the center of the atom

LIVEWORKSHEETS

17. Subatomic particles with a negative charge

Electrons

Quarks

Protons

Neutrons

18. Subatomic particles with a positive charge

neutrons

atomic mass

isotopes

protons

19. Subatomic particles that are neutral in charge

Neutrons

Protons

Electrons

Nucleus

20. A tiny but very dense, positively charged portion of the atom that holds most of the atomic mass

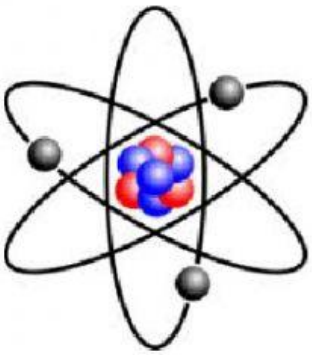
Electron Cloud

Orbitals

Nucleus

Electron Shells

21.



What is one way Neutrons are different from electrons?

Both are in the nucleus(center of atom)

Electrons are inside nucleus and neutrons are outside the nucleus.

neutrons have no charge, while electrons have a negative charge

they both are negative

LIVEWORKSHEETS

22. The mass of this particle is so small that it is not counted.

quark

neutron

electron

proton

23. The term "neutral" means

having only a little bit of charge

having no charge at all

having a negative charge

having a positive charge

24.

Properties of Protons, Neutrons, and Electrons

	Electron	Proton	Neutron
Symbol	e−	p	n
Charge	1−	1+	0
Location	electron cloud around the nucleus	nucleus	nucleus
Relative mass	1/1,840	1	1

True or False - Protons and Neutrons have about the same mass.

False - They do not have about the same mass

True - They have about the same mass

25. Which of the following are found in the nucleus?

protons and neutrons

protons and electrons

Electrons only

protons ONLY

26. What is the charge on a proton?

positive 1

negative 1

0

27. What is the charge on an electron?

0

negative 1

positive 1

LIVEWORKSHEETS

28. What is the mass of a carbon atom that has 6 protons, 7 neutrons, and 6 electrons?

13

7

6

19

29. Which subatomic particle determines the identity of an element?

neutrons

protons

electrons

30. An atom has:
8 Protons
7 Neutrons
6 Electrons
What is the mass number?

LIVEWORKSHEETS