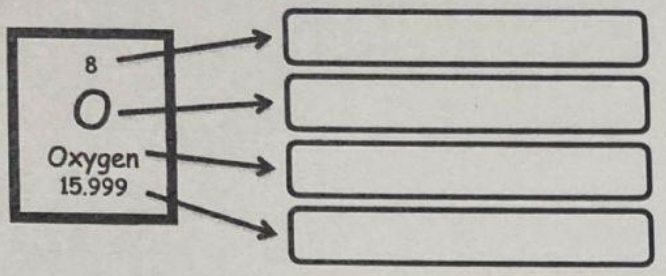
Name: _____ **Date:** _____ **B#:** ____

Atomic Structure:

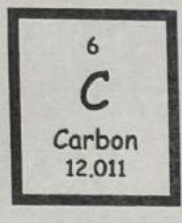


What 2 parts of an atom does the atomic # represent?

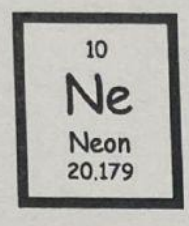
_____ & _____

How do you figure out the # of neutrons?

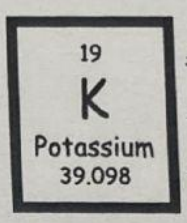
_____



P: ____
E: ____
N: ____



P: ____
E: ____
N: ____



P: ____
E: ____
N: ____

CREATING BOHR DIAGRAMS

Rules for arranging electrons:

1. The 1st energy level can hold up to 2 electrons.
2. The 2nd energy level can hold up to 8 electrons.
3. The 3rd energy level can hold up to 8 electrons.

What term is used to describe the electrons in the outermost energy level?

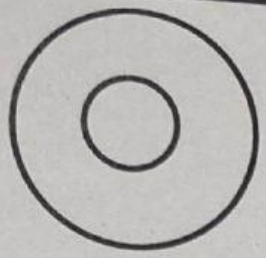
_____

Sketch An Atom	
Draw 5 protons in the nucleus and label with the charge.	
Draw 6 neutrons in the nucleus and label with the charge.	
Draw 2 electrons in the 1st energy level and label with their charge.	
Draw 3 electrons in the 2nd energy level and label with their charge.	
What element is represented?	

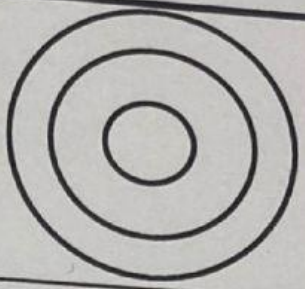
Sketch An Atom	
Draw 3 protons in the nucleus and label with the charge.	
Draw 4 neutrons in the nucleus and label with the charge.	
Draw 2 electrons in the 1st energy level and label with their charge.	
Draw 1 electrons in the 2nd energy level and label with their charge.	
What element is represented?	

LIVEWORKSHEETS

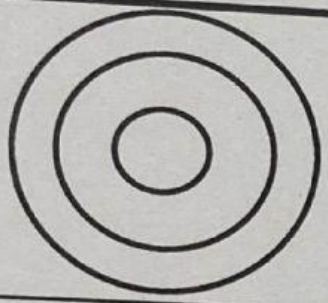
Neon	
# P	
# E	
# N	
# of Valence Electrons	
Atomic #: 10	Mass #: 20



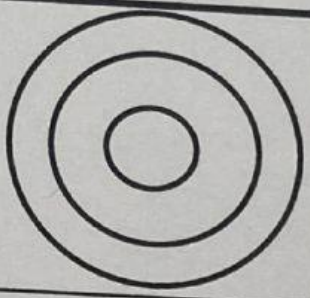
Magnesium	
# P	
# E	
# N	
# of Valence Electrons	
Atomic #: 12	Mass #: 24



Chlorine	
# P	
# E	
# N	
# of Valence Electrons	
Atomic #: 17	Mass #: 35



Silicon	
# P	
# E	
# N	
# of Valence Electrons	
Atomic #: 14	Mass #: 28



CREATING LEWIS DOT DIAGRAMS

Rules for arranging electrons:

1. Figure out how many valence electrons the element has in its atom.
2. Place dots around the element's symbol one at a time (can't exceed 8).

Ne Mg Cl Si

LIVEWORKSHEETS