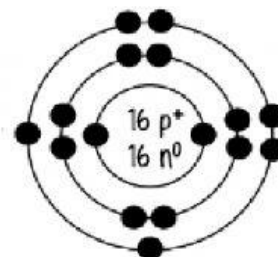


SUM IT UP!

1. This is a Bohr Diagram of a _____ atom.
2. How many energy levels (energy shells) does this atom contain? _____
3. How many electrons does the first energy level contain? _____
4. How many electrons does the second energy level contain? _____
5. How many electrons does the third energy level contain? _____
6. Altogether, how many electrons does this atom have? _____
7. What is the mass number of this atom? _____
8. How many subatomic particles are in the nucleus? _____
9. What is the atomic number of this atom? _____
10. What is the symbol of this atom? _____



Match each statement with the correct word.

11. The smallest particle of an element that retains the properties of that element. _____
12. A positively charged subatomic particle. _____
13. A negatively charged subatomic particle. _____
14. A subatomic particle with no charge. _____
15. The central part of an atom containing protons and neutrons. _____
16. Atoms with the same number of protons but different numbers of neutrons. _____
17. Total number of protons and neutrons in the nucleus. _____
18. The weighted average of the masses of the isotopes of an element. _____

19. Label each part of the Periodic Table Symbol Box.



20. Complete the table below by referencing your own periodic tables. The first one has been completed as an example so you can see how it is to be one.

	Chemical Symbol	Atomic Number	Atomic Mass	Mass Number	Isotopic Notation DO NOT TYPE ANY SPACES	# of protons	# of electrons	# of neutrons JUST TYPE THE NUMBER - DO NOT SHOW WORK
Phosphorous	P	15	30.97	31	Phosphorous - 31	15	15	$31 - 15 = 16$
Aluminium								
Potassium								
Argon								
Lead								

21. Name these atoms with their isotopic name.

