TYPES OF CHEMICAL REACTIONS

DISPLACEMENT (SINGLE AND DOUBLE)

INSTRUCTIONS:

Balance the equations below. Indicate the type of displacement reaction by writing SINGLE or DOUBLE in the space provided. Examples are provided below.

1Cu + **2** $AgNO_3 \rightarrow$ **2** Ag + **1** $Cu(NO_3)_2$ Reaction: **SINGLE** DISPLACEMENT

1 $CuSO_4$ + **2** $NaOH \rightarrow$ **1** Na_2SO_4 + **1** $Cu(OH)_2$ Reaction: **DOUBLE** DISPLACEMENT

QUESTIONS:

1) ___Cl_2 + ___KI →__KCl + __I_2 _____DISPLACEMENT

2) __K_2CO_3 + __$BaCl_2$ → __KCl + __$BaCO_3$ _____DISPLACEMENT

3) ___Na + __$MgCl_2$ → __$NaCl$ + ___Mg _____DISPLACEMENT

4) ___Al + __$CuCl_2$ → ___$AlCl_3$ + ___Cu _____DISPLACEMENT

5) ___Al + ___$Pb(NO_3)_2$ → ___$Al(NO_3)_3$ + ___Pb _____DISPLACEMENT

6) ___$AgNO_3$ + __$CaBr_2$ → __$AgBr$ + __$Ca(NO_3)_2$ _____DISPLACEMENT

7) ___$Al_2(SO_4)_3$ + __$Ca_3(PO_4)_2$ → __$AlPO_4$ + __$CaSO_4$ _____DISPLACEMENT

8) ___$FeBr_2$ + __K_2CO_3 → ___$FeCO_3$ + ___KBr _____DISPLACEMENT

9) ___Cu + ___Ag_2SO_4 → ___Ag + ___$CuSO_4$ _____ DISPLACEMENT

10) ___CuS + ___KCl → ___$CuCl_2$ + ___K_2S _____DISPLACEMENT