

Name: _____ Grade & Section: _____

Worksheet No. _____ **CLASSIFICATION OF MATTER**

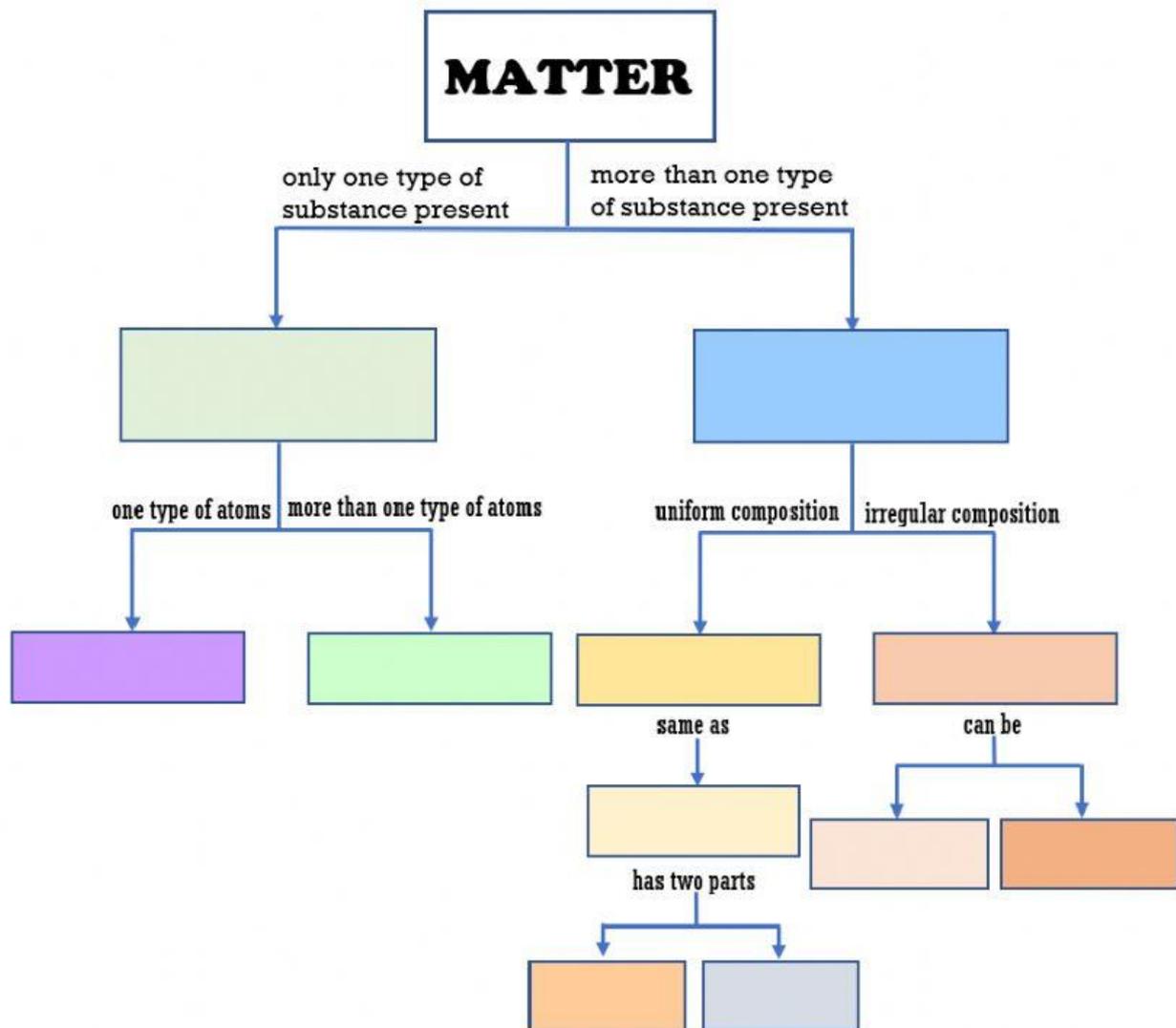
A. Please provide all the boxes with the different classifications of matter. Choose from these given choices:

Solution
Solvent
Mixture

Colloid
Solute
Compound

Suspension
Pure substance
Element

Heterogeneous mixture
Homogeneous mixture
Particle



B. Identify what type of matter is being described. Choose from these given choices:

Solution
Solvent
Mixture

Colloid
Solute
Compound

Suspension
Pure substance
Element

Heterogeneous mixture
Homogeneous mixture
Particle

1. It is a substance that is made entirely from one type of atom. This refers to any substance that cannot be decomposed into simpler substances by ordinary chemical processes.
2. It is a form of matter that has a constant composition and properties that are constant throughout the sample and is composed of one type of atom or molecule.
3. It is the physical combination of two or more elements and/or compounds. It is composed of different types of atoms or molecules that are not chemically bonded.
4. It is a mixture of two or more chemical substances where the various components can be visually distinguished.
5. It is a type of mixture in which the composition is uniform, and every part of the solution has the same properties.
6. It is composed of a particular set of molecules or ions that are chemically bonded.
7. It is defined as a heterogeneous mixture in which the solid particles with a diameter greater than 1000 nm are spread throughout the liquid without dissolving in it. Particles can be separated by standing.
8. It is a homogeneous mixture of substances containing solute and solvent.
9. The component of a mixture that is present in the greatest quantity or that determines the state of matter of the solution and is the one that dissolves a substance.
10. The component of a mixture is present in a lesser quantity, and it is the one being dissolved in a solution.
11. It is a heterogeneous mixture in which microscopically dispersed insoluble particles of one substance are suspended in another substance. The size of the suspended particles in this mixture can range from 1 to 1000 nanometers. This exhibits Tyndall effects.

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