

8Ea

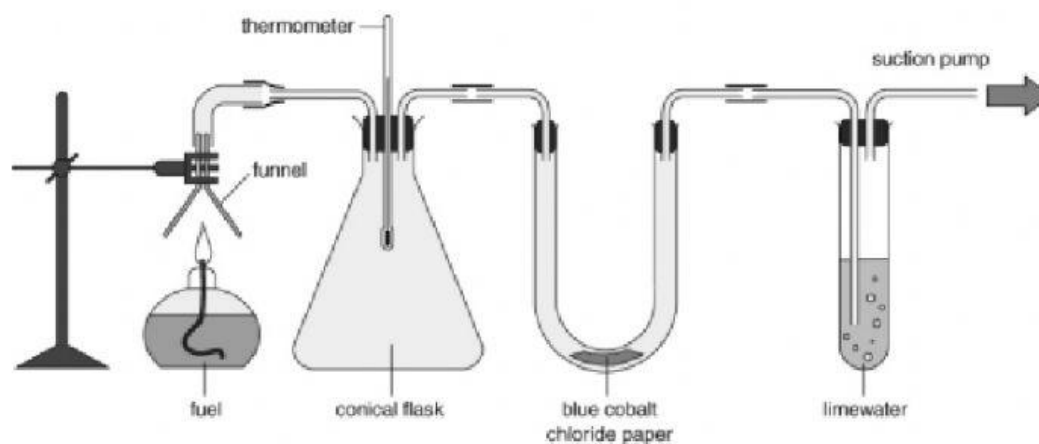
- 1 Combustion of hydrogen produces:
A carbon hydride. **B** hydrocarbon.
C carbon. **D** water.
- 2 Carbon dioxide can be detected using:
A a glowing splint.
B pure water.
C orangewater.
D limewater.
- 3 When hydrated copper sulfate is heated, the gas produced is called:
A the water of crystallisation.
B oxygen.
C carbon dioxide.
D hydrogen.
- 4 methane + oxygen \rightarrow
A hydrogen + oxygen
B carbon dioxide + water
C carbon dioxide + hydrogen
D carbon + water

8Eb

- 1 During a chemical reaction, reactants form products. Compare the mass of reactants with the mass of products formed.
A The mass of reactants is greater than the mass of products.
B The mass of products is greater than the mass of reactants.
C The mass of products is the same as the mass of reactants.
D The change in mass depends on what the reactants and products are.

- 2 Oxidation is always:
- A a reaction in which a substance combines with oxygen.
 - B a reaction in which a substance burns in oxygen.
 - C a reaction in which a substance burns in air.
 - D an explosive reaction.
- 3 When zinc burns in oxygen it forms:
- A water.
 - B zinc hydroxide.
 - C zinc carbonate.
 - D zinc oxide.
- 4 When 4.8 g of magnesium reacts with 3.2 g of oxygen, how much magnesium oxide is formed?
- A 1.6 g
 - B 4.9 g
 - C 7.0 g
 - D 8.0 g

Below is a diagram showing tests used for hydrocarbon combustion.



a What does the cobalt chloride paper test for? _____

b What does the limewater test for? _____

c How will the thermometer reading change during the experiment?

d Explain why the thermometer reading changes in this way.

Complete the following word equations by dragging and dropping the correct molecules.

Oxygen	Hydrogen	Carbon dioxide	Iron
Copper	Water	Magnesium oxide	Carbon

a _____ + Oxygen \longrightarrow Carbon dioxide

b Magnesium + Oxygen \longrightarrow _____

c Methane (hydrocarbon) + _____ \longrightarrow _____ + _____

d _____ + Oxygen \longrightarrow Copper oxide

e _____ + Oxygen \longrightarrow Water

f _____ + Oxygen \longrightarrow Iron oxide