DIRECTIONS: FOR EACH PROBLEM. DRAG THE VALUES INTO THE CORRECT PLACES. THEN, CALCULATE THE ANSWER (*round to 3 decimal places*)

1. How many grams are in 6.3 moles of sulfur? *32.07 g* *6.3 moles* *1 mole* *6.02 x 10^{23} atoms*

$$\frac{\square}{1} \times \frac{\square}{\square} = \square$$

2. How many moles are in 7 g of carbon? *7 g* *1 mole* *12.01 g* *6.02 x 10^{23} atoms*

$$\frac{\square}{1} \times \frac{\square}{\square} = \square$$

3. How many atoms are in 7.1 g of sodium? *6.02 x 10^{23} atoms* *7.1 g* *1mole* *22.99 g*

$$\frac{\square}{1} \times \frac{\square}{\square} = \square$$

4. How many atoms are in 0.35 moles of aluminum? *6.02 x 10^{23} atoms* *1 mole* *0.35 moles* *26.98 g*

$$\frac{\square}{1} \times \frac{\square}{\square} = \square$$

5. How many moles are in 22,000,000 atoms of calcium? *22,000,000 atoms* *6.02 x 10^{23} atoms* *1 mole* *40.08 g*

$$\frac{\square}{1} \times \frac{\square}{\square} = \square$$

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