Sitat of electrolyte and non-electrolyte

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Digital student worksheets (LKPD)

student name. :

Class :

No. Absence. :

Instruction:

Read and understand the following material:

Electrical conductivity in solution

A solution is a homogeneous mixture of two or more substances solution. Composed of a solvent and a solute. Based on the electrical conductivity, the nature of the solution can be influenced by the type of substance dissolved in a solution. Substances that can dissolve in water are divided into electtolytes and non-electrolytes. This difference is based on the presence of electrical conductivity in the solution.

K-3 Understand, apply, and analyze factual, conceptual, and procedural knowledge, based on their curiosity about science,technology,art,culture, and humanities with insight into humanity, nationality, state, and civilization, related to the causes of phenomena and events in a specific field of work to solve the problem

K-4 rocessing, reasoning, and presenting in the concrete and abstract realms are related to the development of what they learn in school independently, and being able to carry out specific tasks under direct supervision.





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KD	KD FORMULA
KD. 3.5	Analyzing the properties of electrolyte and non-electrolyte sulotions
KD. 4.5	Distinguish between the propoties of electrolyte and non-electrolyte solutios.

Competency Achievement indicators (IPK)

CODE IPK	IPK formulation
3.5.1	Analyzing the symptoms of electric current conduction in a solution
3.5.2	Desinguish between electrolyte and non-elektrolyte solutions
3.5.3	Identify electrolytes and non-electrolytes
3.5.4	Classify electrolytes and non-electrolitytes
3.5.5	Distinguish between electrolyte and non-electrolyte solution based on current conduction electricity.

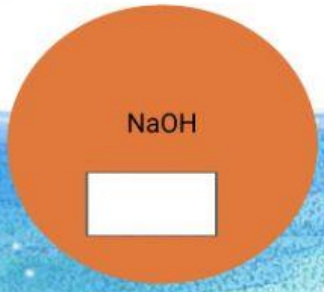
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Purpose:

Thourg the discovery learning model, student are to analyze the state of the hsolition based on its electrical conductivity with The character of curiosity, active inteligent, and honest and can develope critical thinking skills, vommunicate, collaborate creativily (4C) in distinguishing the electrical conductivity of various solutions through the design and axecution of experiments.

1. Following solution are classified as electrolyte solution
   A. Alkohol
   B. On coal
   C. Kitchen salt water
   D. Urea
   E. Glukosa
2. The following pairs of solution are classified as non- electrolytes...
   A. Urea and drinking water
   B. Urea and sugar
   C. Acetic acid and ammonia
   D. Table salt and sulfuric acid
   E. Sea water and table salt
3. Name the following chemical formula!





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4. Find a matching pair by drawing a line!

Ammonium hydroxide	H2SO4
Sulfurid acid	NaCl
Sodium cloride	NHaCH



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