Name of the student: Teacher: Sridhar Sriram

Class/Section:

Subject: Science (Chemistry)

Chemical bonding

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1. Lesson Story: A is a lasting attraction between atoms, ions or molecules that enables the formation of chemical compounds. The bond may result from the electrostatic force of attraction between oppositely charged ions as in or through the sharing of electrons as in . The strength of chemical bonds varies considerably; there are "strong bonds" or "primary bonds" such as covalent, ionic and metallic bonds, and "weak bonds" or "secondary bonds" such as dipole–dipole interactions, the London dispersion force and hydrogen bonding.

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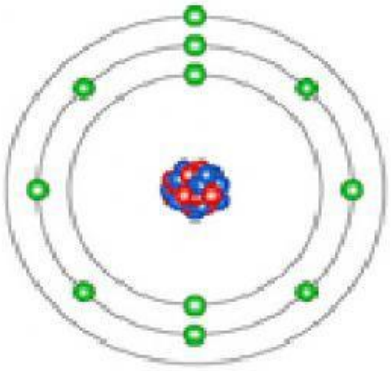
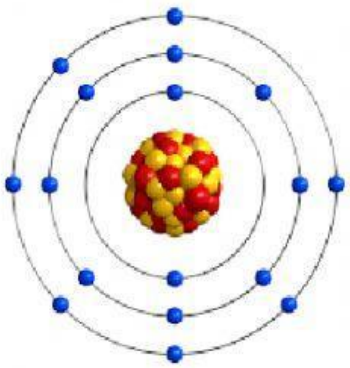
2. Find the subatomic particles for the given atoms and write the number of electrons needed to form the bonds. Identify the type of bonds between the atoms.



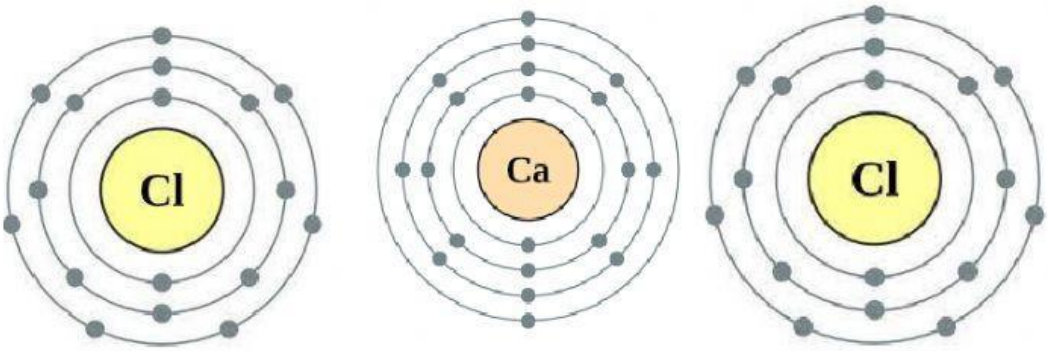


Element

Element

Symbol	Symbol
Atomic number	Atomic number
Total electrons	Total electrons
Number of shells	Number of shells
Metal/Nonmetal	Metal/Nonmetal
Valence electrons	Valence electrons
Lose /Gain electrons	Lose / Gain electrons
Type of bond	Type of bond

3.Find the subatomic particles for the given atoms and write the number of electrons needed to form the bonds. Identify the type of bonds between the atoms



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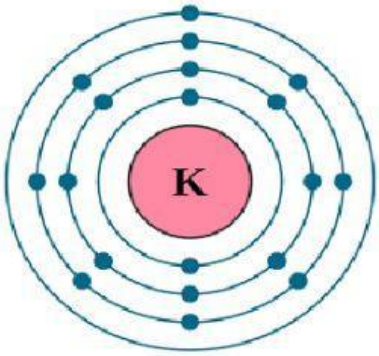

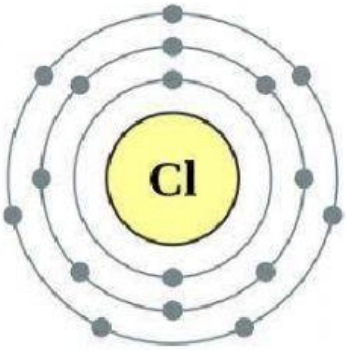

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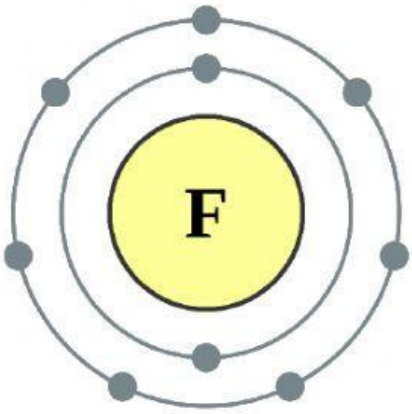

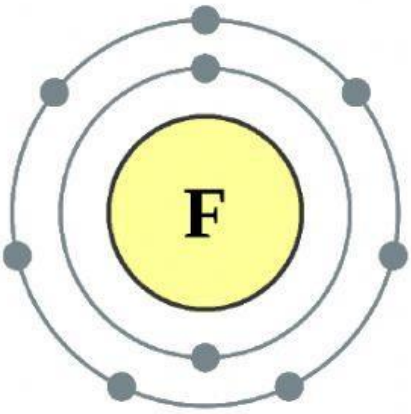

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Lose /Gain/Share	Lose / Gain/Share
Type of bond	Type of bond

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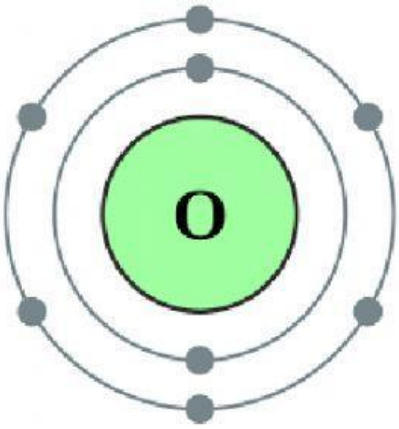

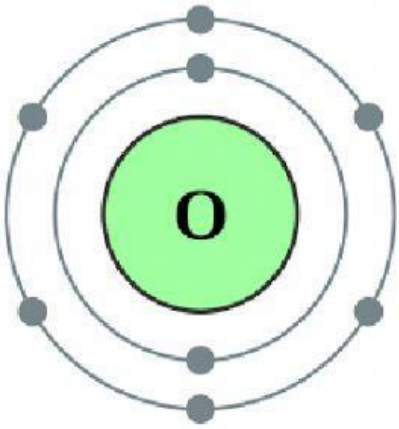

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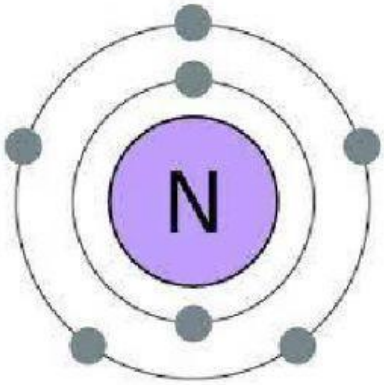

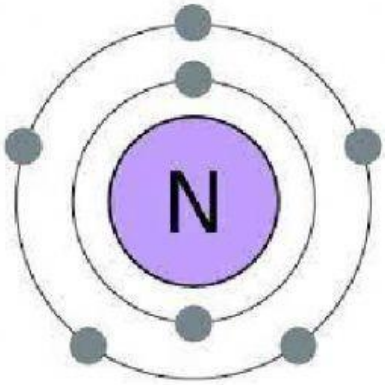

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