

1+	1	
1	2+	Oxidation Number
1 H Hydrogen	-2 2	Valence Electrons
		Family
3 Li Lithium	4 Be Boron	
11 Na Sodium	12 Mg Magnesium	
19 K Potassium	20 Ca Calcium	

Oxidation Number	3+	4+/-	3-	2-	1-	0
Valence Electrons	3	4	5	6	7	8
Family						
5 B Boron	6 C Carbon	7 N Nitrogen	8 O Oxygen	9 F Fluorine	10 Ne Neon	
13 Al Aluminum	14 Si Silicon	15 P Phosphorus	16 S Sulfur	17 Cl Chlorine	18 Ar Argon	
31 Ga Gallium	32 Ge Germanium	33 As Arsenic	34 Se Selenium	35 Br Bromine	36 Kr Krypton	

Some Polyatomic Ions and their Oxidation Numbers				
1+	1-	2-	3-	
ammonium (NH ₄)	acetate (C ₂ H ₃ O ₂)	carbonate (CO ₃)	phosphate (PO ₄)	
	chlorate (ClO ₃)	sulfate (SO ₄)		
	hydroxide (OH)			
	nitrate (NO ₃)			
	bicarbonate(HNO ₃)			

	Symbols & Oxidation #	Formula	Name of Compound
15. Magnesium and chlorine			
16. Potassium and acetate			
17. Aluminum and sulfate			
18. Lithium and nitrogen			

19. An ion is an atom or group of atoms that has become electrically _____.

20. When an atom loses an electron its charge is (**positive or negative**)

21. An ionic bond is the attraction between (**opposites, positive, neutral, or negative**) ions.

22. Ionic compounds are electrically (**charged, positive, neutral, or negative**).

23. The sum of the charges for an ionic compound is _____.

The two answers must be in the right order.

24. An ionic compound is the result of the bonding of a (**non-metal, metalloid, metal, noble gas**) with a (**non-metal, metalloid, metal**).