Exercise Electronic configuration of Manganese

1) **Write the electronic configuration of Manganese at its ground state:**

$_{25}Mn$: 1s 2s 2p 3s 3p 4s 3d

2) How many valence e belong to the **Manganese**:

3) Determine the outermost orbital in **Manganese**:

4) **Write the electronic configuration of Mn ion after 2 e are donated:**

$_{25}Mn$: 1s 2s 2p 3s 3p 4s 3d

5) **Write the electronic configuration of Mn ion after 4 e are donated:**

$_{25}Mn$: 1s 2s 2p 3s 3p 4s 3d

Give a set of quantum numbers for the e in the lowest energy orbital:

1st e: n = , l = , m= , s=

or

2nd e: n = , l = , m= , s=

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