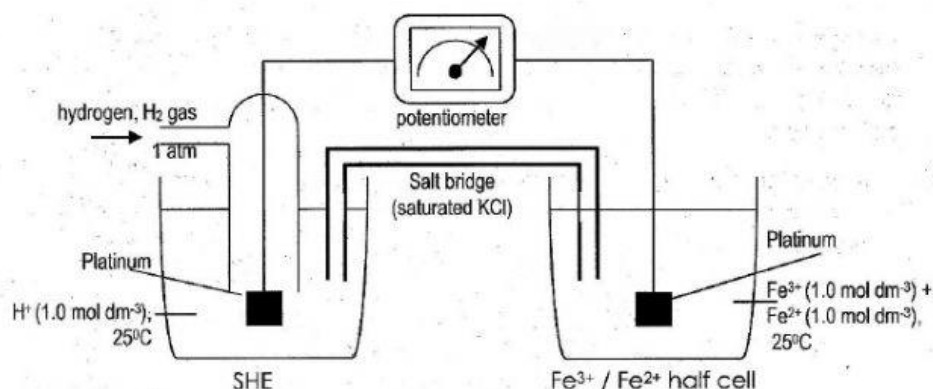


Fill in the blanks



$$E^{\circ} \text{ for } \text{H}^+/\text{H}_2 = 0\text{V}, E^{\circ} \text{ for } \text{Fe}^{3+}/\text{Fe}^{2+} = +0.77\text{V}$$

Oxidation occur at _____ and reduction occur at the _____. During the redox reaction, H_2 is _____ to H^+ and the oxidation number _____ from 0 to +1. While Fe^{3+} is _____ to Fe^{2+} and the oxidation number _____ from +3 to +2. The oxidising agent is _____ and the reducing agent is _____.

The observation is _____ solution turn into _____ solution

decrease	increase	H_2	oxidised	reduced
$\text{Fe}^{3+}/\text{Fe}^{2+}$ half cell	green	Fe^{3+}	SHE	yellow