CHEMBUDDY CHAPTER 3

3.1 CLASSIFICATION OF ELEMENTS



CHOOSE THE CORRECT ANSWER

NO	QUESTION	ANSWER
1	What is group in periodic table? A. The horizontal row with label 1 to 7 B. The horizontal row with label 1 to 18 C. The vertical columns with label 1 to 18 D. The vertical columns with label 1 to 7 only	A B C D
2	What is period in periodic table? A. The horizontal row with label 1 to 7 B. The horizontal row with label 1 to 18 C. The vertical columns with label 1 to 18 D. The vertical columns with label 1 to 7 only	A B C D
3	Which block is NOT TRUE in periodic table? A. b block B. d block C. f block D. p block	A B C D
4	Deduce period of element Y in periodic table. Given the electronic configuration of Y as $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 4s^2\ 3d^2$ A. period 1 B. period 2 C. period 3 D. period 4	A B C D
5	What is the charge on all the simple ions of elements of group 17? A. 1+ B. 2+ C. 1- D. 2-	A B C D
6	Sodium (Z=11) and Silicon (Z=14) are examples of two elements which belong to the same A. group B. period C. block D. compound	A B C D

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7	$^{37}_{17}P$ What is the position of this element in the periodic table? Block / Group / Period A. p / 17 / 3 B. p / 15 / 3 C. p / 7 / 3 D. s / 5 / 3	A B C D
8	$^{63}_{29}Q$ What is the position of this element in the periodic table? Block / Group / Period A. s / 2 / 4 B. s / 12 / 4 C. d / 9 / 3 D. d / 11 / 4	A B C D
9	An element R has a valence electronic configuration of $5s^2\ 4d^6$. What is the period and block of R in the Periodic Table? Block / Period A. s / 5 B. s / 4 C. d / 5 D. d / 4	A B C D
10	An electronic configuration for element T is $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 4s^2\ 3d^{10}\ 4p^4$. Determine the group of element T A. 4 B. 14 C. 6 D. 16	A B C D
11	An electronic configuration for element U is $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 4s^2\ 3d^{10}\ 4p^6\ 5s^2\ 4d^8$. Determine the group of element U. A. 8 B. 10 C. 14 D. 16	A B C D

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12	The proton number for element V is 32. What is the valence orbital of element V? A. 3d 4p B. 4p C. 4s 4p D. 3d	A B C D
13	Elements in the Periodic Table are arranged in order increasing A. nucleon number B. proton number C. number of electron D. number of neutron	A B C D
14	Elements in the same group will have similar A. conductor properties B. thermal properties C. physical properties D. chemical properties	A B C D
15	Elements in Group 2 are known as A. alkaline earth metal B. noble gases C. alkali metal D. halogen	A B C D



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