21. The charge of an ion P is 2-, in which it contains 2 inner electrons and eight outermost electrons. The electronic configuration of atom P is;

A. $1s^2\ 2s^2\ 2p^6$

B. $1s^2\ 2s^2\ 2p^4$

C. $1s^2\ 2s^2\ 2p^6\ 3s^2$

D. $1s^2\ 2s^2\ 2p^6\ 3s^2$

22. Which of the following electronic configuration represent an element that form ion with a charge of 2-.

A. $1s^2\ 2s^2\ 2p^6\ 3s^2$

B. $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^2$

C. $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^4$

D. $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^5$

23. Choose the most suitable reason of the anomality in electronic configuration of Chromium with the proton number of 24.

A. Stability of fully filled 3d orbital.

B. Stability of half-filled orbital.

C. Stability of half-filled 4s orbital.

D. Stability of half-filled 3d orbital.

24. Determine the electronic configuration of the most stable ion of element X-25.;

A. $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 4s^2\ 3d^5$

B. $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 4s^2\ 3d^6$

C. $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 3d^5$

D. $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 4s^2\ 3d^{10}$

25. Shown below are a set of quantum number of the highest energy electron in P^+ ion. Determine the electronic configuration of P atom.

n=4, l=0, m=0, s= +½

A. $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 4s^1$

B. $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 4s^2$

C. $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6$

D. $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 4s^2\ 3d^1$