

21. The charge of an ion P is $2-$, in which it contains 2 inner electrons and eight outermost electrons. The electronic configuration of atom P is;
- A. $1s^2 2s^2 2p^6$
B. $1s^2 2s^2 2p^4$
C. $1s^2 2s^2 2p^6 3s^2$
D. $1s^2 2s^2 2p^6 3s^2$
22. Which of the following electronic configuration represent an element that form ion with a charge of $2-$.
- A. $1s^2 2s^2 2p^6 3s^2$
B. $1s^2 2s^2 2p^6 3s^2 3p^2$
C. $1s^2 2s^2 2p^6 3s^2 3p^4$
D. $1s^2 2s^2 2p^6 3s^2 3p^5$
23. Choose the most suitable reason of the anomaly in electronic configuration of Chromium with the proton number of 24.
- A. Stability of fully filled 3d orbital.
B. Stability of half-filled orbital.
C. Stability of half-filled 4s orbital.
D. Stability of half-filled 3d orbital.
24. Determine the electronic configuration of the most stable ion of element X-25.;
- A. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^5$
B. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
C. $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5$
D. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10}$
25. Shown below are a set of quantum number of the highest energy electron in P^+ ion. Determine the electronic configuration of P atom.
- $n=4, l=0, m=0, s=+\frac{1}{2}$
- A. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$
B. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$
C. $1s^2 2s^2 2p^6 3s^2 3p^6$
D. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^1$