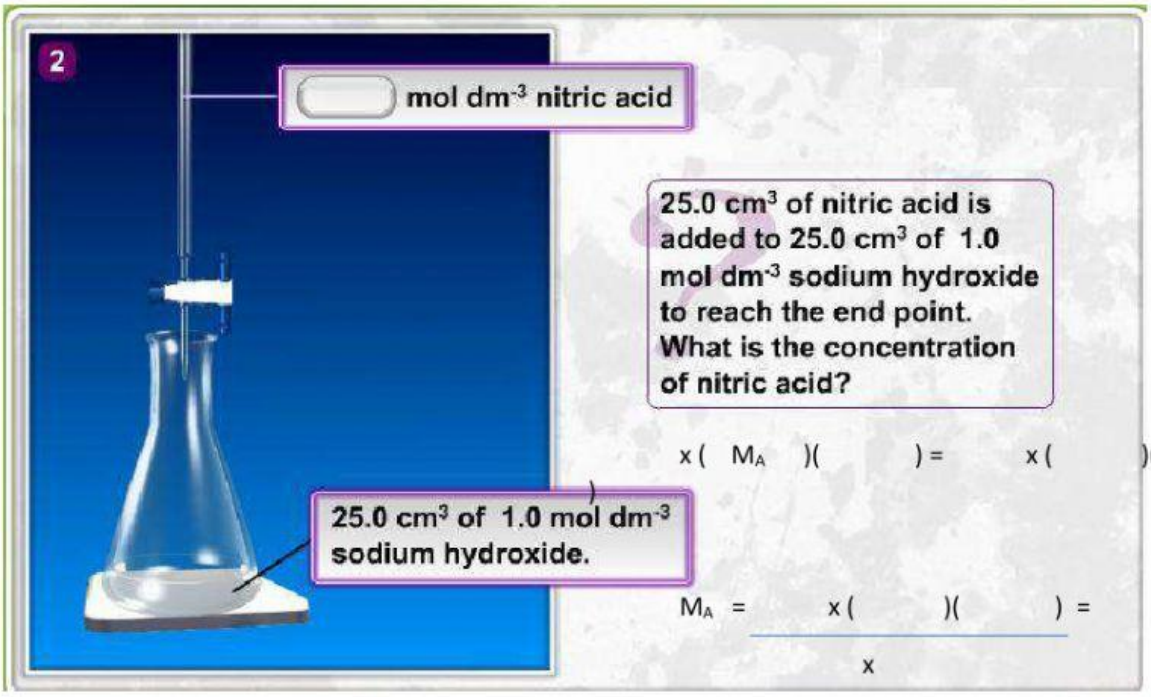
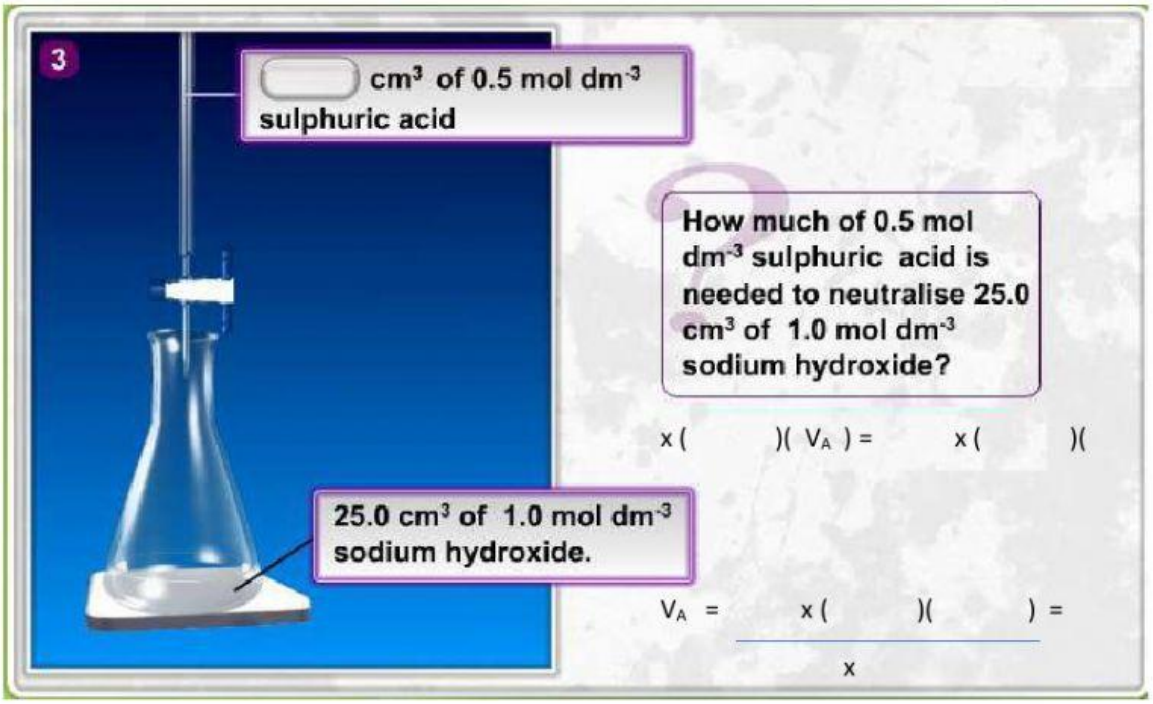
CALCULATION OF NEUTRALISATION THROUGH TITRATION METHOD

Question 1



Question 2



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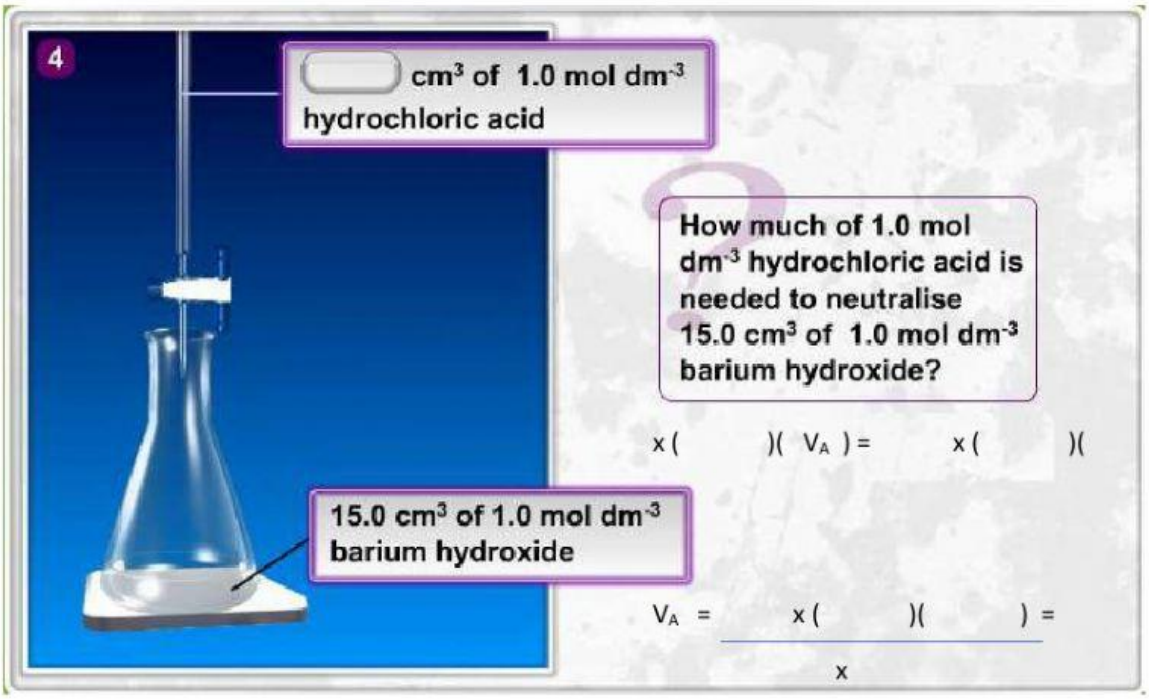
Question 3



Question 4

If it takes 25 mL of 0.05 M HCl to neutralize 345 mL of NaOH solution, what is the concentration of the NaOH solution?

x (M_B)( ) = x ( )( )

M_B = x ( )( ) = moldm^{-3}
x

LIVEWORKSHEETS

Question 5

If it takes 50.0 mL of 0.5 M $Ca(OH)_2$ solution to completely neutralize 125 mL of HNO_3solution, what is the concentration of the HNO_3 solution?

x (M_A)( ) = x ( )( )

M_A = x ( )( ) = $moldm^{-3}$

x

Question 6

A 25.0 mL sample of HCl was titrated to the endpoint with 15.0 mL of 2.0 M NaOH. What is the molarity of HCl in gdm^{-3}? [RAM: H= 1, Cl = 35.5]

x (M_A)( ) = x ( )( )

M_A = x ( )( ) =

x

Change to gdm^{-3}: x = gdm^{-3}

Question 7

If it takes 75.0 mL of 1.5 M HNO_3 solution to completely neutralize 125 mL of $Ca(OH)_2$what is the concentration of the $Ca(OH)_2$ solution in gdm^{-3} ? [RAM : Ca = 40 , O = 16, H = 1]

x (M_B)( ) = x ( )( )

M_B = x ( )( ) =

x

Change to gdm^{-3}: x = gdm^{-3}

LIVEWORKSHEETS