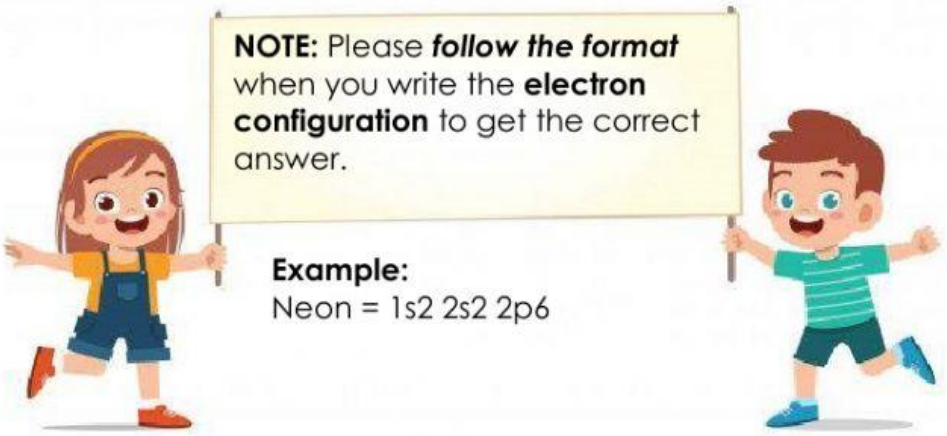
Quiz No. 1
Electron Configuration

I. Do the following.

A. Write the complete *electron configuration* for each atom.
B. Give the Highest Energy Level of the electron
C. Give the valence electron
D. Write your answers inside the box.



1. Argon **Electron Configuration** –

Highest Energy Level –
Valence Electron –

2. Sulfur **Electron Configuration** –

Highest Energy Level –
Valence Electron –

3. Radium **Electron Configuration** –

Highest Energy Level –
Valence Electron –

II. Identify what elements are represented by each of the following electron configurations? **(4 pts)**

a. $1s^2\ 2s^2\ 2p^5$
b. $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 4s^2$
c. $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 4s^2\ 3d^{10}4p^5$
d. $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 4s^2\ 3d^{10}\ 4p^6 5s^2\ 4d^{10}\ 5p^4$

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NOTE: Please **follow the format** when you write the ***orbital configuration*** to get the correct answer.

Example:
Neon = 1s2 2s2 2px2 2py2 2pz2



III. Write the ***orbital configuration*** of the following elements.

a. Aluminum –

b. Selenium –

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