

NAME: _____

DATE: _____

Let's start
calculate the
Molarity!

EXPERIMENT 2: DATA ANALYSIS



READ THE QUESTION FIRST.

A titration of 25.00 mL of an x M HCl solution with 0.15 M NaOH starts at a burette reading for NaOH of 0.20 mL. The burette reading of the end point is 24.10 mL.



How much is
NaOH used
in this exp?

Now calculate the no. of moles
of NaOH dispensed.

Volume of NaOH dispensed

=

=

Moles of NaOH = MV

=

$\times$

=

We need to
know how
many
moles of
NaOH so
we can
determine
no. of moles
of HCl
reacted.

Let's write the
balanced
equation for the
neutralisation
reaction.

Balanced equation:

+



+

Calculate the number of HCl present in the acid solution. (DO stoichiometry)

react with

react with

Formula

Molarity = _____

Substitution & Final answer

=

x =

